



**K.L.E. SOCIETY'S
BASAVAPRABHU KORE ARTS, SCIENCE AND COMMERCE
COLLEGE, CHIKODI – 591 201.**

Accredited at A* Grade With 3.42 CGPA in 4th Cycle

Website: klesbkcollegechikodi.edu.in ☎ : 08338 – 272176 Email: kles_bkcc@rediffmail.com

2.2.1 Assessing the learning level of students and organizing programmes:

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K.L.E. SOCIETY'S EST-1969
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DATE: 20/12/2023

DEPARTMENT OF ENGLISH

Induction Test 2023-24

TALENT LEVEL ASSESSMENT TEST (TLAT)

For English Opt. Students at Entry Level

Time: 30 Minutes

Max. Marks: 25

Name: Ms. Tejaswini. Dundage

class: BA-I

Roll No: 60

*Encircle the Correct Answer and fill in the blanks All the Questions are Compulsory.

- 1) Who wrote the play *Romeo and Juliet*?
a) John Lyly b) Shakespeare c) Thomas Kyd
- 2) How many consonants are there in English?
 a) 21 b) 26 c) 5
- 3) I will go to College. (Name the tense of this sentence)
a) Past b) Present c) Future
- 4) _____ is definite article. a) the
- 5) Seeta ate an apple (identify the article)
a) Seeta b) an c) apple
- 6) Give synonym for the word **ugly**.
 a) Bad b) Cute c) Beautiful
- 7) Give antonym for the word **hard**.
a) Delete b) Destroy c) Soft
- 8) The medicinal creeper has small leaves resembling betel leaves. (leaves/lives)
- 9) Lucia looked very weak. (weak/week) Who wrote Discovery of India?
a) Kuvempu b) Jawaharlal Nehru c) Tagore
- 10) Who has written '*The Gentleman of the Jungle*'?
a) B. C. Chatterjee b) Jomo kenyatta c) Kautilya

11) Who is the author of *Wings of Fire*?

a) APJ Abdul Kalam

b) R. K. Narayan

c) Kalidas

12) Find out the hyphen mark.

a) ?

b) :

c) -

13) Who wrote *My Experiments with Truth*?

a) Shakespeare

b) Girish Karnad

c) M. K. Gandhi

14) Who wrote our National Anthem?

a) Kuvempu

b) Anantamurthy

c) Tagore

15) Who wrote 'Vandemataram'?

a) Bankimachandra Chatterjee

b) Dr. D. R. Bendre

c) Tagore

16) Who is called the 'Shakespeare' of India?

a) Kalidas

b) R. K. Narayan

c) Rajarao

17) Who is known as 'Nightingale of India'?

a) Lata Mangeskar

b) Sarojini Naidu

c) M. S. Subbalaxmi

18) Who wrote Sonnet 'True Love'?

a) Girish Karnad

b) Tolstoy

c) Shakespeare

19) How many parts of speech are there in English?

a) 7

b) 8

c) 6

20) Mark the pronoun.

a) you

b) for

c) from

21) Sonnet is a _____ line poem.

a) 16

b) 14

c) 18

22) Who is the founder of 'Shantiniketan'?

a) Kuvempu

b) Nehru

c) Tagore

23) Give prefix for the word **happy**.

a) Un

b) in

c) dis

24) I am a liar, _____?

a) Did I?

b) am I?


c) aren't?

25) Usage of the Wh-word **Where**.

a) Place

b) Time

c) Person


Sign of the student

82.5%



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Date: 06/11/2023

DEPARTMENT OF BOTANY
INDUCTION TEST 2023-24

Class: B.Sc. I Semester

Date: 6/11/2023

Name: *Umaefomkh. L. Burakhe*

Total Marks : 20M

Roll No: 28

Marks Obtained: 13

Read the following questions and mark the correct answer.

1. Who is father of Biology?
a. Aristotle ~~b. Theophrastus~~ c. Robert Hooke d. Ramachandran
2. Which smallest cell among the following?
~~a. Bacteria~~ b. Virus c. Mycoplasma d. Fungi
3. cell without cell wall is called...
a. cell b. cytoplasm ~~c. protoplast~~ d. nucleus
4. membrane of vacuole is called as
a. chloroplast b. centriole c. tonoplast d. ER
5. Which of the following is called as power house of cell
a. Ribosome b. Nucleus c. DNA ~~d. mitochondria~~
6. Who discovered a living cell
~~a. Engler~~ b. Watson ~~c. Robert hooke~~ d. Anton von leewenhok
7. charge of DNA is...
~~a. Neutral~~ b. +ve ~~c. -ve~~ d. all of the above
8. Choose abiotic factor from the following
~~a. goat~~ b. bacteria c. tree ~~d. water~~

- 2 ✓ 9. Length of Human DNA is.....
a. 2 meters b. 2.1 meters ~~c. 2.2 meters~~ d. 2.3 meters
- X 10.is the basic unit of classification and a taxonomic rank.
a. species ~~b. Genus~~ c. class d. order
- ✓ 11. Who discovered penicillin?
a. Robert brown b. Watson c. mendel ~~d. Alexander Fleming~~
- ✓ 12. Genetic material in prokaryotes is called _____
~~a. Nucleoid~~ b. Nucleus c. Nucleolus d. Nuclei
- ✓ 13. Which of the following group of organisms are called as "Amphibians of plant kingdom"
a. Algae b. Bryophytes ~~c. Pteridophytes~~ d. Gymnosperms
- ✓ 14. Water conducting tissue in plants is....
a. Meristem b. Parenchyma c. Phloem ~~d. Xylem~~
- ✓ 15. Bacteria belong to which of the following kingdom?
a. Mycota b. Plantae ~~c. Monera~~ d. Protista
- ✓ 16. Fungal cell wall is made up of...
a. Pectin ~~b. chitin~~ c. cellulose d. lignin
- ✓ 17. Double fertilization and triple fusion is the characteristic of....
a. Algae b. Bryophytes ~~c. Angiosperms~~ d. gymnosperms
18. Which of the following plant is selected by Mendel for experiment
a. Solanum nigrum b. Pisum sativum c. Mangifera indica d. Basella alba
- ✓ 19. Phenotypic ratio of monohybrid cross is.....
a. 1:3 ~~b. 3:1~~ c. 2:1 d. 3:1
- X 20. Who is called as syndrella of genetics?
~~a. Honey bee~~ b. Horse c. Drosophila melanogaster d. Butterfly



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DEPARTMENT OF ENGLISH

Details of Slow and Advanced Learners

2023-24

Sl. No.	Roll No.	Name of the Student	PU Marks	% PU	Specific Subject Marks at PU	TLAT Marks	Average %	Slow / Advanced
1	01	SHRUSHTI V KANDAGANVE	511	85.16	82	16 / 64	77.05	AL
2	05	LAXMI SANJU MALAGE	546	91	74	20 / 80	81.66	AL
3	12	RESHMA S NAIK	521	86.83	67	16 / 64	72.61	AL
4	25	SHIVANAND R GHARABUDE	-	-	-	18 / 72	-	-
5	26	AJIT PARAPPA PATIL	-	-	-	-	-	-
6	27	SAKSHI RAMESH BANSODE	493	82.16	54	20 / 80	72.05	AL
7	34	POOJA RAJU MALAGE	554	92.33	73	22 / 88	84.44	AL
8	35	MUKTABAI B DASAR	531	88.5	66	17 / 68	74	AL

9	53	ROHINI RAMESH KAMBLE	448	74.66	56	17 / 68	66.22	SL
10	57	AISHWARYA C DODDAMANI	478	79.66	50	18 / 72	67.22	SL
11	60	TEJASWINI D DUNDAGE	357	59.5	73	21 / 84	72.16	AL
12	70	SURAKSHA G HIREMATH	439	73.16	56	19 / 76	68.38	SL
13	72	SAPNA S KAPALI	334	55.66	40	19 / 76	57.22	SL
14	73	SEEMAH GUJALAR	519	86.5	58	17 / 68	70.83	AL
Out of 14 students, 04 are slow learners and 08 are advanced learners						Class	71.98	SL & /AL
02 are absent.						Average		



Teacher in-Charge



Head, Dept. of English
Head

Department of English



Principal

PRINCIPAL

K. S. Basavaprabhu College
Arts, Science and Commerce
CHIKODI - 591 207



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DEPARTMENT OF BOTANY

INDUCTION TEST DSCC 2023-24

Marks Sheet					
Sl.No.	Roll No.	Name	Marks Obtained (20)	II P.U.C. %	Average
1	17	Radhika A. chougala	5	78.78	41.89
2	28	Umarfaruk L Burakhe	13	82.5	47.75
3	34	Bharatesh L. Kamate	14	78	46
4	48	Pallavi S. Badode	15	84	49.5
5	50	Anupama A. Patil	13	71	42
6	53	Kamarunissa A. Sayyad	15	73.5	44.25
7	56	Madhuchandrika R. Sadalage	12	58.4	35.2
8	76	Sakshi S Kalal	15	76.95	45.975
9	82	Kamble vishwajeet Anil	6	52	29
10	84	Spoorti G Dangi	10	89.66	49.83
11	85	Suprema S Jogalekar	12	74	43
12	89	Bhoomika S Chavan	14	77.83	45.915
13	90	Kavita S. Bendigeri	10	67.33	38.665
14	93	Shweta B Hundekar	18	89.83	53.915
15	95	Basagouda S Koti	9	58	33.5
16	97	Gayatri M Mariyai	17	81.6	49.3
17	104	Kaveri L. Boraganvi	6	83.16	44.58
18	108	Prajwal P. Gadage	15	58	36.5
19	109	Vijay A. Mujawar	14	53.68	33.84
20	111	Sudharani K Gudodagi	15	80.83	47.915
Class Average			12.4	73.4525	42.92625

Class Average	42.92
Total students appeared	20
Advanced learners	13
slow learners	7

Advanced learners-13

Sl.No.	Roll No.	Name	Marks Obtained (20)	H.P.U.C. %	Average
1	93	Shweta B Hundekar	18	89.83	53.915
2	84	Spoorti G Dangi	10	89.66	49.83
3	48	Pallavi S. Badode	15	84	49.5
4	97	Gayatri M Marivar	17	81.6	49.3
5	111	Sudharani K Gudodagi	15	80.83	47.915
6	28	Umarfaruk L. Burakhe	13	82.5	47.75
7	34	Bharatesh L. Kamate	14	78	46
8	76	Sakshi S Kalal	15	76.95	45.975
9	89	Bhoomika S Chavan	14	77.83	45.915
10	104	Kaveri L. Boraganvi	6	83.16	44.58
11	53	Kamarunissa A. Sayyad	15	73.5	44.25
12	85	Suprema S Jogalekar	12	74	43
13	50	Anupama A. Patil	13	71	42

Slow learners-07

Sl.No.	Roll No.	Name	Marks Obtained (20)	H.P.U.C. %	Average
1	17	Radhika A. chougala	5	78.78	41.89
2	90	Kavita S. Bendigeri	10	67.33	38.665
3	108	Prajwal P. Gadage	15	58	36.5
4	56	Madhuchandrika R. Sadalage	12	58.4	35.2
5	109	Vijay A. Mujawar	14	53.68	33.84
6	95	Basagouda S Koti	9	58	33.5
7	82	Kamble vishwajeet Anil	6	52	29


HOD Botany

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**MATHEMATICS DEPARTMENT
PRACTICE FOR SLOW AND ADVANCE LEARNERS
2023-2024**


Students have to think critically in all respective courses. Each student has their own thinking/IQ level. Courses which are required critical thinking always aim at helping students to improve their skills and also their IQ level through their learning and social activities. After admission at entry level, the department assesses the learning levels of the students and classifies the students as advance learner and slow learners. Later the department plan to organize special activities that can be conducted for advanced learners and slow learners. After admission to B.Sc first semester in Oct/Nov 2023, the details of the students noted down. We have categorized the students as slow and advance learners on the basis of

- 1) Second PUC marks and
- 2) Marks scored in mathematics subject
- 3) Marks scored in Induction Test.

The average of these three indicators is prepared. The students who are secured less than 70% are considered as slow learners and above 70% are considered as advance learners.

Some activities are conducted for advanced learners such as conducting class seminars/discussions, advised them to solve previous year question papers, guided for their future career and encourage them to participate in the various competitions: quiz, seminars and workshops.

For slow learners we advised them to practice some important theorems and guided them to solve the previous years question papers.


HOD. D.
MATHEMATICS
B. K. College, CHIKODI




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DEPARTMENT OF MATHEMATICS
Slow and Advance Learner
(Year 2023-24)

Roll No.	Registration No.	Name of the Student	II PUC %	II PUC Marks	II PUC Maths Marks	TLAT Marks	Average %	Remark
1	U15DM23S0001	SHIVANI PARAMANAND PATIL	82.50	495	85	13	80.56	AL
4	U15DM23S0004	BHAVESHWARI NINGOUDA PATIL	64.83	389	45	14	60.28	SL
7	U15DM23S0007	ABHISHEK SHANKAR KAMBAR	75.00	450	70	AB	72.22	AL
8	U15DM23S0008	RAHUL RAJU PUJARI	54.83	329	40	AB	51.25	SL
9	U15DM23S0009	KEERTI MAHADEV GUDODAGI	70.50	423	73	12	68.89	SL
10	U15DM23S0010	KALIKA SAMBHAJIRAO SHENDURE	54.67	328	35	8	50.42	SL
11	U15DM23S0011	SAMIKSHA RAGHUNATH JADHAV	72.00	432	57	8	67.92	SL
14	U15DM23S0014	AKSHATA APPASAB LATTE	63.00	378	52	13	59.72	SL
16	U15DM23S0016	LAXMI BASAVARAJ KOKATANUR	82.00	492	87	14	80.42	AL
18	U15DM23S0018	RAHUL SHIVANAND SAMPAGAR	66.67	400	66	AB	64.72	SL
21	U15DM23S0021	AISHWARYA AJIT TORASE	87.17	523	96	12	85.97	AL
22	U15DM23S0022	SHIVARAJ DUNDAPPA KOTABAGI	95.00	570	98	13	92.78	AL
24	U15DM23S0024	SHRUSTI SURESH BIKKANAVAR	61.83	371	59	14	59.72	SL
26	U15DM23S0026	ARPITA KEMPANNA HOLEPPAGOL	74.83	449	80	11	73.47	AL
27	U15DM23S0027	VIRUPAKSHA RAMESH DATTAWADE	51.00	306	67	12	51.81	SL
30	U15DM23S0030	SANGEETA BHIMAPPA CHOUGALA	83.50	501	90	12	82.08	AL
33	U15DM23S0033	ZIYA JAVEED MULLA	85.17	511	90	AB	83.47	AL
36	U15DM23S0036	LAXMI SHIVAKUMAR KOLI	74.00	444	71	14	71.53	AL

37	U15DM23S0037	MAHASHWETA MAHADEV BHISTE	74.50	447	78	6	72.92	AL
41	U15DM23S0041	SHANKAR VITTAL BHAGAWATI	86.67	520	95	12	85.42	AL
43	U15DM23S0043	SHREYA BALAPPA VANTAGUDE	71.67	430	65	8	68.75	SL
45	U15DM23S0045	ROHIT BABASAB KAMBLE	58.33	350	50	AB	55.56	SL
49	U15DM23S0049	CHAITALI SADASHIV KHOT	77.83	467	84	9	76.53	AL
51	U15DM23S0051	SINDHU MAHANTESH NOOLI	73.83	443	69	14	71.11	AL
60	U15DM23S0060	KAVERI SANJU CHANDAKI	61.50	369	48	11	57.92	SL
62	U15DM23S0062	SANIKA SHIVANAND MARATHE	63.83	383	66	AB	62.36	SL
64	U15DM23S0064	JYOTI THAKKANAVAR	84.33	506	94	14	83.33	AL
65	U15DM23S0065	SUKANYA SHREEKANT BALOL	83.17	499	81	12	80.56	AL
66	U15DM23S0066	GOUTAMI RAJU INGALE	87.67	526	97	13	86.53	AL
71	U15DM23S0071	AISHWARYA MAYUR HAVALDAR	80.17	481	87	13	78.89	AL
72	U15DM23S0072	SANDEEP GOPAL GUDASI	76.33	458	75	8	74.03	AL
73	U15DM23S0073	TEJASHWINI AJEET JODATTI	84.00	504	93	9	82.92	AL
74	U15DM23S0074	MEGHA DUNDAPPA HALINGALI	96.17	577	98	12	93.75	AL
75	U15DM23S0075	SANIKA SHIVANAND SWAMI	75.00	450	71	11	72.36	AL
77	U15DM23S0077	JYOTI KALLAPPA ICHALAKARANJI	86.17	517	91	12	84.44	AL
78	U15DM23S0078	MAHEK FAIROZ PATEL	73.83	443	55	14	69.17	SL
80	U15DM23S0080	PANKAJ GANAPATI MANE	75.67	454	70	12	72.78	AL
83	U15DM23S0083	SAGAR KUMAR PATIL	90.83	545	96	9	89.03	AL
86	U15DM23S0086	AISHWARYA LAKKAPPA BIDARI	83.50	501	95	12	82.78	AL
87	U15DM23S0087	SANIYA ABASAHEB PATEL	54.83	329	30	14	49.86	SL
94	U15DM23S0094	SHILPA RAJU MUTARE	83.00	498	84	12	80.83	AL
98	U15DM23S0098	POOJA DAYANAND HIREMATH	87.00	522	96	11	85.83	AL
99	U15DM23S0099	SAMPADA YASHAVANT KUMBAR	89.33	536	90	12	86.94	AL
101	U15DM23S0101	SWATI SANJAY HINDALKAR	85.33	512	93	13	84.03	AL
103	U15DM23S0103	PRIYA LAXMANGOURA PATIL	93.67	562	99	12	91.81	AL
105	U15DM23S0105	ANKITA SADASHIV MARENNAVAR	81.67	490	86	9	80.00	AL
106	U15DM23S0106	SHIVANI BANAKARE	79.50	477	93	11	79.17	AL
107	U15DM23S0107	PALLAVI SHIVABASU SHIVAPUR	71.67	430	64	10	68.61	SL

110	U15DM23S0110	ANKITA ISHWAR PATIL	86.00	516	91	10	84.31	AL
114	U15DM23S0114	CHAITRA KRISHNA KAMBAR	82.50	495	73	9	78.89	AL

AL-Advance learner & SL-Slow learner


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K.L.E Society's
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COLLEGE CHIKODI

NAME : PRIYANKA KANDAGAYE

SUB : STRATEGIC COST MANAGEMENT

R.NO : 59

CLASS : 6th sem (B.COM)

REG NO : C2024857

Submitted to.

DR. LAXMIKANTH SIR.

1) Limitations of Traditional cost management

1. It ignore competition market growth and customer requirements
2. Traditional cost management excessively focus on cost reduction
3. A broad based cost reduction could lead to low quality of products and services
4. which might drive away customer resulting in lower sales and profitability
4. Excessive focus on cost reduction could impact on quality of product and services and negatively impact on customer.
5. There is a limited focus on review and improvisation.
6. It has in short term out looks
7. Saving cost on an annual bases
8. It ignore development marketing and economics

Needs of strategic cost management

- 1) It is cost analysis in a broad context where strategic eliminates become more explicit and formal strengthening strategic position of an company.
- 2) Cost data analysis and used strategically to develop alternate major to gaining sustainable competitive advantages
- 3) Strategic cost management is give a clear understanding of the companies cost structure in such of sustainable competitive advantage.

4) Strategic cost management is a managerial review of cost information explicitly directed to the four stages of SCM (formulation, communication, implementation and control).

5) Strategic cost management helps in overall recognition of cost relationships among the activities in value chain.

6) Explain the advantages and disadvantages of JTI

- Right quantity at right time

2) Higher quality - the approval of large batch manufacturing and reduction in handling often results in significant quality improvements often in the region of 25% or more.

3) Improve Customer Service

4) Minimal inventory and waste: one of the main aims with any JIT implementation is to improve stock turns and the amount of stock being held. Personal experience has seen reduction of more than 90% stock in some industries, along with the reduction in the stock came many other associated benefits.

Reduced lead time: one of the most significantly impacted areas is that of the time it takes for products to flow through the process. Instead of weeks or months most JIT implementations result in lead

of hours or a few days

6) Increase equipment utilization

7) simplicity - the use of simple pull systems such as Kanban, even with your suppliers, can significantly reduce the need for any form of complex planning. with many implementations result: the only planning is the final shipping process.

8) more workforce participation - one requirement of JIT as with most other aspects of lean manufacturing is that employees are heavily involved in the design and application of your system.

9) continuous improvement

10) Better problem solving

11) Reduce cost

Just in Time Limitations :-

- * Requires high degree of delegation
- * Requires change in philosophy and culture
- * vulnerability in break down
- * work only in regular products
- * Requires highly reliable & flexible suppliers

3) Explain the Advantages & disadvantages of BPR (Business process reengineering)

1) The main advantage of business process reengineering (BPR) is to identify any processes, sub processes, costs, and labour that are not required in an organisation. It not only highlights the deficiencies of a procedure but also shows area where business is incurring costs without getting required returns.

2) The BPR provides management an opportunity to understand the systems of their business in more detail. It improves the communication within the enterprise and also enhances the quality of internal management reports.

3) The BPR approach provides an insight of business analysis to management which can help in future decision making as it presents a better picture of the strengths and weakness of the business and business's ability to adapt and respond towards the changes in external environment.

4) BPR provides long term solutions to minimize the deficiencies of a business. The focus of the exercise is to take short term damage control measures but specifically initiate a long term plan enabling for improvement in efficiency, capability and effectiveness of the organisation.

Disadvantages of BPR

- ① Although the idea of process reengineering is easy to understand in a literal manner, its practical implementation is very difficult. BPR process reengineering not only requires monetary assets of a company but also consumes time and human resources.
- 2 Business process reengineering may not suit to every business. The size, availability of sources & needs of every business are different. BPR usually benefits large organizations.
- 3 The approach does not provide immediate results. Incremental improvements in processes can be seen instantly as these focus upon the costs of business process but BPR focuses majorly upon long term revenue synergies of a business which take time to crystallize and are difficult to forecast.
- 4 The sustainability of a change in a business process is a difficult task. BPR follows many aspects that are unfavorable for certain levels of business i.e. redundancies, cost cutting changes etc. So, the development of a consensus among all the stakeholders of the organization to experience and continue the processes after change has occurred is difficult.

15 marks

- ① ABC manufacturing sales its product at ₹1000 per unit due to competition, its competitors are likely to reduce price by 15%. B wants to respond aggressively by cutting price by 20% and expects that the present volume of 150000 units p.a will increase 200000. B wants earn 10% target profit on sales based on the following.

particulars	existing	Target cost
Direct material cost per unit	400	385
Direct manufacturing labour	55	50
Direct machinery cost p.u	70	60
Direct manufacturing cost per unit	525	495

Other manufacturing oh

no. of orders (80 per order)	22500	21250
testing hours (5.2 per hour)	4050000	30,00,000
units reworked (2100 PU)	12000	13000

manufacturing ohs are also allocated using relevant cost drivers. Other operating cost per unit for the expected volume are

estimated as follows

Research & Design	50
marketing & customer service	130
Total →	180

- 1) calculate target cost per unit & target cost for the proposed volume showing breakup of a different cost elements.
- 2) prepare targeted profitability statement.

particulars	units	amt
present selling price	1000	
(-) R. price Reduction cost		
[1000 x 20%]	2000	
target profit $800 \times \frac{10}{100}$	800	
	80	
target cost	720	

statement showing break-up target cost per unit

material cost	385	7,70,00,000
labour cost	50	1,00,00,000
machinery cost	60	1,20,00,000
add: manufacturing oh		
order cost (212.50 x 80)	8.5	17,00,000
Testing hours (30,00,000 x 2)	30.	60,00,000
Reworked cost (13,000 x 100)	6.50	13,00,000
manufacturing cost	540	10,80,00,000
add: other operating cost		
Research & Design (50 x 200,000)	50	10,00,00,000
marketing [130 x 2,00,000]	13.0	26,00,00,000
target cost	720	14,40,00,000

3) Star, Co. Ltd currently only 10,000 units of product X which is selling at ₹ 80 per unit the cost of producing product X ₹ 75 per unit production can be increased 10,500 by utilizing idle capacities provided additional output can be sold in in the market on the bases of market research the company introducing target costing cost of product X can be brought down substantially through new design for a product and changes in the next process of manufacturing the estimate for next year or target selling price ₹ 70 per unit target sales volume 10000 units
 Target profit margin 10% on sales
 calculate target cost per unit and target cost of expected volume and compare existing profit with target profit

Computation of Target Cost per unit

particulars	target cost p.u	target cost value
		₹ 12000
Selling price	70	8,40,000
(-) profit 10% on sale	7	84,000
Target cost	63	7,56,000

computation with existing cost with target cost

	existing	10,000 units	target	10,500 units
S.p	80 pu	800,000	70 pu	735,000
(-) cost	75 pu	750,000	63 pu	661,500
profit	5	50,000	7	73,500

2) ABC Company is planning to buy a new machine for ₹ 7,00,000. The company is confident for selling all the units produce with the help of this machine to yield annual cash surplus of ₹ 200,000 but there is some uncertainty about the working life of machine. However it is found that 250 machines (of this modern) are used at present by other companies. These companies informed the following. Useful life pattern further @ 10% discount rate and net present value of each different machine life is presented below.

A. machine life in year	3	4	5	6	7
No. of machines	20	50	100	70	10 = 250
N.p.v	2026000	660000	582000	710000	2736000

Advice about purchase of machine

particulars	3	4	5	6	7	Total
	20	50	100	70	10	
	250	250	250	250	250	
	= 0.08	= 0.2	0.4	0.28	0.04	
Npv	2026000	660000	582000	710000	2736000	
(P x NPV)						
(Expected value)	162080	132000	232800	478800	109440	526960

3) The cost structure of product m manufactured by ABC company is $5x + 50,00,00,000$ where x represents yearly output product m has 4 year market life and during this period the company expect to produce & sell 5,00,00,000 units 10,00,00,000 units 15,00,00,000 units 20,00,00,000 units 25,00,00,000 units respectively in the year 1 to 5 you are required to determine:

- year wise unit cost of product m under conventional costing.
- life cycle cost per unit under product life cycle costing

variable cost = 5 times of output
 Total cost = $5x + 50,00,00,000$ (fixed cost)
 x = yearly output

unit cost statement

year	out put crore	var cost	fixed cost crore	Total cost ₹ crore	unit cost
1	5 crore	$5 \times 5 = 25$	50	75	$75/5 = 15$
2	10	$(5 \times 10) = 50$	50	100	$100/10 = 10$
3	15	$5 \times 15 = 75$	50	125	$125/15 = 8.33$
4	20	$5 \times 20 = 100$	50	150	$150/20 = 7.5$
5	25	$5 \times 25 = 125$	50	175	$175/25 = 7$
Total	75	375		625	7.83

PART I

Answer any ten of questions

(10x2=20)

1. (a) Define Fourier transformation.

If $f(x)$ is periodic function of x , then the Fourier integral of $f(x)$ may be expressed as,

$$f(x) = \frac{1}{2\pi} \int_{-\infty}^{+\infty} e^{i\omega x} dt \int_{-\infty}^{+\infty} f(t) e^{-i\omega t} dt = \frac{1}{2\pi} \int_{-\infty}^{+\infty} e^{i\omega x} F(\omega) d\omega$$

Where

$$F(\omega) = \frac{1}{2\pi} \int_{-\infty}^{+\infty} f(t) e^{-i\omega t} dt$$

The function $F(\omega)$ is called the Fourier transform of $f(t)$.

1. (b) Find the Laplace transformation of $L\{e^{at}\}$.

Solution:

$$L\{f(t)\} = \int_0^{\infty} e^{-st} f(t) dt$$

put $f(t) = e^{at}$

$$\begin{aligned} L\{e^{at}\} &= \int_0^{\infty} e^{-st} e^{at} dt = \int_0^{\infty} e^{-st+at} dt = \int_0^{\infty} e^{-(s-a)t} dt = \left[-\frac{e^{-(s-a)t}}{(s-a)} \right]_0^{\infty} \\ &= \left[-\frac{1}{(s-a)} \times e^{-(s-a)\infty} \right] - \left[-\frac{1}{s-a} \times e^{-(s-a)0} \right] \\ &= 0 + \frac{1}{(s-a)} e^0 \quad \because e^0 = 1 \\ &= \frac{1}{(s-a)} \end{aligned}$$

1. (c) State Parseval's identities of Fourier transform.

$$i) \int_{-\infty}^{+\infty} F(\omega) \bar{G}(\omega) d\omega = \int_{-\infty}^{+\infty} f(t) \bar{g}(t) dt$$

Where $\bar{G}(\omega)$ =Complex conjugate of $G(\omega)$

$\bar{g}(t)$ =Complex conjugate of $g(t)$

$$ii) \int_{-\infty}^{\infty} [F(\omega)]^2 d\omega = \int_{-\infty}^{\infty} [f(t)]^2 dt$$

1. (d) What is PIN diode? *p-type and n-type of highly doped material separated by lightly doped intrinsic region.*

A PIN diode is simply a p-n junction diode that is designed to have a very small junction capacitance (0.01 to 0.1 pf) or A PIN diode is a semiconductor device that operates as a variable resistor at RF and microwave frequencies.

1. (e) Mention the application of Optical fibers.

Applications of Optical fibers

1. Telecommunication
2. LAN and WAN
3. Factory Automation
4. Premises Wiring
5. Medical Lab

1 (f) What is meant by virtual ground in op-amp?

Virtual ground: In op-amplifier the potential difference across the input impedance is zero and if any one terminal is grounded then the potential at the other end is zero. It is called virtual ground.

1. (g) Write the syntax of if-else statement?

C if...else statement

The if...else statement is used if the programmer wants to execute some statement/s when the test expression is true and execute some other statement/s if the test expression is false.

Syntax of if...else

if (test expression)

{ Statements to be executed if test expression is true;

}

Else

{ Statements to be executed if test expression is false;

1 (h) Write any two application of IC- 555 timer

Application of IC- 555 timer

1. Can be used as Monostable or Astable Multivibrator.
2. Used for DC-DC Converters.
3. Used in Digital Logic Probes.
4. Used as Waveform generators.
5. Used in Temperature measurements and control devices.
6. Used for Burglar and Toxic Gas alarms.
7. Used as Voltage Regulators.
8. Used to create some time delays.

1 (i) Write a C-programme to print "Good Morning"

```
#include <stdio.h>

int main()
{
    printf("Good Morning \n");
    return 0;
}
```

1 (j) What is break statement? Where it is used.

Break statement

Terminates the **loop** or **switch** statement and transfers execution to the statement immediately following the loop or switch.

Uses of Break Statement

In C programming, break is used in terminating the loop immediately after it is encountered. The break statement is used with conditional if statement.

1 (k) Calculate the modulation factor for AM wave if $V_{\max}=4V$ and $V_{\min}=2V$

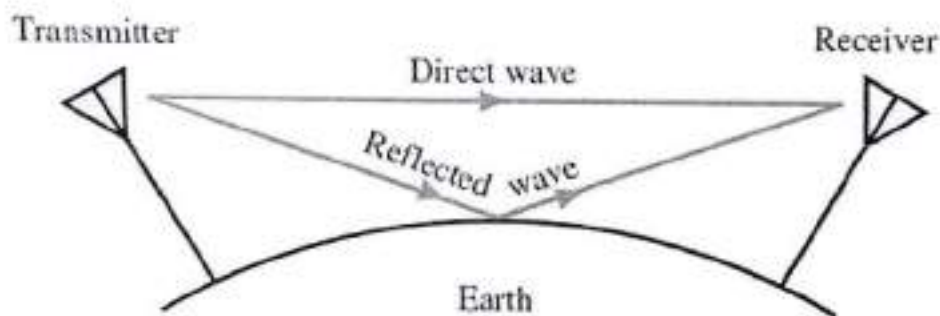
$$m = \frac{V_{\max} - V_{\min}}{V_{\max} + V_{\min}} = \frac{4 - 2}{4 + 2} = \frac{2}{6} = 0.3333 \times 100 = 33.33\%$$

4	N.Aperature	Numerical aperture remains the same	NA changes continuously with distance from fiber axis
5	Materials used	Normally plastic or glass is preferred.	Only glass is preferred.
6	Attenuation	Typically 0.34 dB/km at 1.3 μm	More 0.6 to 1 dB/km at 1.3 μm
7	Applications	In local network communication	In LAN and WAN

4. Explain Space wave propagation.

Space wave propagation

The space wave propagation is that made of propagation in which the space waves travel in a straight line from transmitting antenna to receiving antenna. Space waves are radio waves of frequency range from 54MHz to 4.2GHz. Space waves are used for line of sight communication (LOS) as well as for the satellite communication. The space waves is used for television broad cast, microwaves link and satellite communication.



The communication through space wave between transmitter and receiver is limited to line of sight (LOS) path. This line of sight communication is limited by

- (1) The line of sight distance and
- (2) The curvature of the earth

The space waves following the line of sight propagation can be blocked at some points by the curvature of the earth. If the signal is to be received beyond the horizon then the receiving antenna must high enough to intercept the line of sight waves.

If the transmitting antenna is a height h_t it can be shown that the distance to the horizon is $d_t = \sqrt{2Rh_t}$. Where R is the radius of the earth. d_t is also called radius horizon of the transmitting antenna. The maximum line of sight distance d_M between the transmitting antenna and the receiving antenna is

$$dM = \sqrt{2Rh_T} + \sqrt{2Rh_R}$$



Where h_T is the height of transmitting antenna and h_R is the height of receiving antenna. dM is called the range of communication.

5 Write a C-program to convert the temperature Centigrade to Fahrenheit using the relation $F = 32 + \frac{9}{5}C$.

*/*Converting Centigrade to Fahrenheit using the relation $F = 32 + \frac{9}{5}C$.*/*

```
# include <stdio.h>
# include <conio.h>
void main()
{
float c, f;
clrscr();
printf("Enter the Centigrade : ");
scanf("%f", &C);
F = [ 32+(9/5) * C ];
printf("\n The Fahrenheit is : %.2f", F);
getch ();

}
```

6. An AM wave is represented by expression: $V = 5(1 + 0.8\cos 10^4 t)\sin 220 \times 10^4 t$ volts.

(i) Find the minimum and maximum amplitudes.

(ii) Find the frequency components contained in the modulated wave and the amplitude of each component.

The instantaneous voltage of AM wave is given by $V_m = V_a(1 + m\cos \omega_m t)\cos \omega_c t$

The given Amplitude Modulated wave is $V = 5(1 + 0.8\cos 10^4 t)\sin 220 \times 10^4 t$

(i) The maximum amplitude of AM wave

$$V_{m(\max)} = V_a(1 + m) = 5(1 + 0.8) = 5(1.8) = 9 \text{ Volts}$$

The minimum amplitude of AM wave

$$V_{m(\min)} = V_s(1 - m) = 5(1 - 0.8) = 5(0.2) = 1 \text{ Volts}$$

Frequency components of the wave

(i) Carrier wave $f_c = \frac{\omega_c}{2\pi} = \frac{220 \times 10^4}{2\pi} = 350 \text{ KHz}$

(ii) Signal wave $f_s = \frac{\omega_s}{2\pi} = \frac{10^4}{2\pi} = 1.59 \text{ KHz}$

(ii) Frequency of components contained in a modulated wave

(i) Carrier wave $f_c = 350 \text{ KHz}$

(ii) Lower sideband $f_c - f_s = (350 - 1.59) \text{ KHz} = 348.41 \text{ KHz}$

(iii) Upper sideband $f_c + f_s = (350 + 1.59) \text{ KHz} = 351.9 \text{ KHz}$

Amplitude of three components of modulated wave

(i) Carrier wave $V_c = 5 \text{ volts}$.

(ii) Lower sideband $\frac{mV_c}{2} = \frac{0.8 \times 5}{2} = 2 \text{ volts}$.

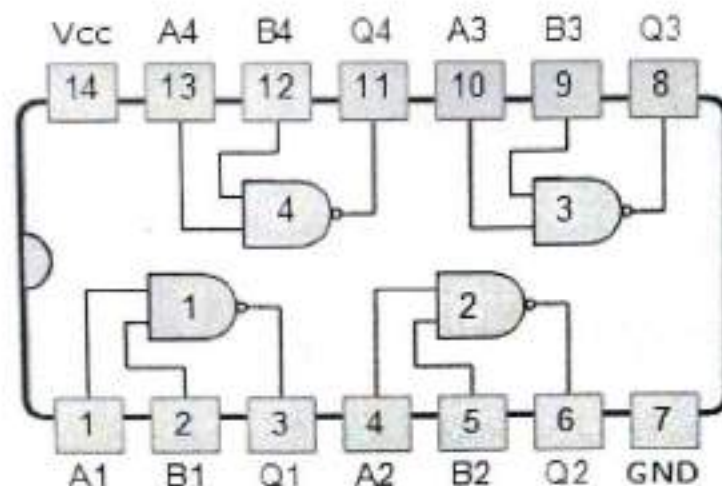
(iii) Upper sideband $\frac{mV_c}{2} = \frac{0.8 \times 5}{2} = 2 \text{ volts}$.

7. Explain the function of various pins of IC-7400.

The IC 7400

The 7400 IC package contains four independent positive logic NAND GATES with open-collector outputs. Pins 14 and 7 provide power for all four logic gates. The outputs require pull-up resistors to function properly. The open-collector outputs can be connected together to produce higher output voltages or to implement an active-LOW wired-OR function or active-high wired-AND function without the need for additional logic gates. In Boolean algebra any Boolean equation can be represented with NAND gates alone.

IC-7400



- Pin 1:** This is the input terminal for first NAND gate.
Pin 2: This is the input terminal for first NAND gate.
Pin 3: This is the output terminal for first NAND gate
Pin 4: This is the input terminal for second NAND gate.
Pin 5: This is the input terminal for second NAND gate.
Pin 6: This is the output terminal for second NAND gate
Pin 7: Ground Terminal: All voltages are measured with respect to this terminal.
Pin 8: This is the output terminal for third NAND gate.
Pin 9: This is the input terminal for third NAND gate.
Pin 10: This is the input terminal for third NAND gate.
Pin 11: This is the output terminal for fourth NAND gate.
Pin 12: This is the input terminal for fourth NAND gate.
Pin 13: This is the input terminal for fourth NAND gate.
Pin 14: Supply Terminal: Supply voltage is applied to this terminal with respect to ground.

PART III

Answer any four of the following

(4x10=40)

8. (i) State and explain shifting property of Fourier transform.

Shifting property of Fourier transform

If $F(\omega)$ is the Fourier transform of $f(t)$: the Fourier transform $f(t \pm a)$ will be given by $e^{\pm i\omega a} F(\omega)$
 Where a is any constant.

Proof:

By definition of finite Fourier transform

$$F.T\{f(t \pm a)\} = \frac{1}{\sqrt{2\pi}} \int_{-\infty}^{+\infty} f(t \pm a) e^{-i\omega t} dt$$

Substituting $(t \pm a) = y$
 i.e. $dt = dy$ we have,

$$F.T\{f(t \pm a)\} = \frac{1}{\sqrt{2\pi}} \int_{-\infty}^{+\infty} f(t \pm a) e^{-i\omega(y \mp a)} dy \quad \because t \pm a = y, \therefore t = y \mp a$$

$$F.T\{f(t \pm a)\} = e^{\pm i\omega a} \frac{1}{\sqrt{2\pi}} \int_{-\infty}^{+\infty} f(y) e^{-i\omega y} dy \quad \because i\omega y - i\omega(\mp a) = 0, \quad -i\omega y \pm i\omega a = 0$$

10 (ii) Derive Expression for frequency modulation

Expression for frequency modulation

The general form of sinusoidal voltage equation is represented by

$$v = V \cos \theta(t) \text{-----(a)}$$

The time rate change of $\theta(t)$ gives the angular frequency of the voltage,

$$\omega = \frac{d\theta}{dt} \text{-----(b)}$$

For the unmodulated carrier wave the angular frequency ω_c is constant. Hence in this case we get

$$\theta(t) = \int \omega_c dt = \omega_c t + \phi \text{-----(c)}$$

The modulating signal can be written as

$$v_m = V_m \cos \omega_m t \text{-----(d)}$$

Where V_m is amplitude of signal and ω_s is frequency of signal voltage.

The general form of unmodulated carrier voltage equation is represented by

$$v_c = V_c \sin (\omega_m t + \theta) \text{-----(e)}$$

Where $\phi(t) = (\omega_m t + \theta)$ is the total instantaneous 'phase' of function

So $v_c = V_c \sin \phi(t) \text{-----(f)}$

The instantaneous angular frequency is $\omega_i = \frac{\partial \phi(t)}{\partial t} = \frac{\partial}{\partial t} (\omega_c t + \theta) = \omega_c \text{-----(g)}$

By definition, the instantaneous angular frequency of modulated wave can be written as

$$\omega_i = \omega_c + K_f V_s \cos \omega_m t \text{-----(h)}$$

Similarly the frequency of the unmodulated carrier is f_c then the instantaneous value of the carrier frequency of the modulated wave is given by

$$f_i = f_c + K_f V_m \text{-----(h)}$$

Where K_f is constant called the frequency deviation constant. Substituting for v_m its sinusoidal value we get,

$$f_i = f_c + K_f V_m \cos \omega_m t \text{-----(i)}$$

Fig shows the variation of f_i with time 't'. In fig (b) amplitude of carrier remains constant.

From this the maximum variation of the carrier frequency from its unmodulated value is

$$\Delta f = K_f V_m \text{-----(j)}$$

This is called peak frequency deviation.

The total variation in frequency from the lowest to the highest is called the carrier swing. The carrier swing is obviously equal to double the frequency deviation

From the above equation (i) &(j) we have

$$f_i = f_c + \Delta f \cos \omega_m t \text{----- (k)}$$

The frequency deviation $\Delta f = f_{max} - f_c = \frac{K_f V_m}{2\pi}$ or $\Delta f = f_c - f_{min} = \frac{K_f V_m}{2\pi}$

Multiplying on both side by 2π to the above equation (k) we get angular frequency

$$\omega_i = \omega_c + \Delta \omega \cos \omega_m t \text{----- (l)}$$

The modulation index is the ratio of frequency deviation to the modulated frequency.

$$m = \frac{\text{frequency deviation}}{\text{modulated frequency}} = \frac{\Delta f}{f_m} = \frac{K_f V_m}{\omega_m}$$

Therefore the expression for FM wave becomes

$$v = V_c \cos(\omega_c t + m \sin \omega_s t)$$

11 (i) Explain while statement and for statement with syntax.

C programming while and do...while Loop Syntax

While (test expression)

{ Statement/s to be executed.

}

The while loops checks whether the test expression is true or not. If it is true, code/s inside the body of while loop is executed, that is, code/s inside the braces { } are executed. Then again the test expression is checked whether test expression is true or not. This process continues until the test expression becomes false.

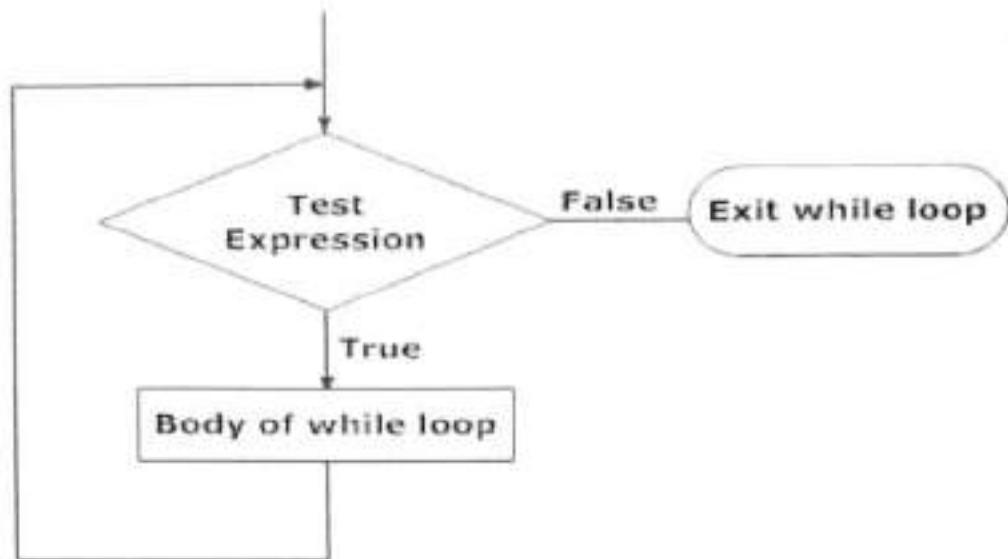


Figure: Flowchart of while loop

C Programming for Loop

Loops cause program to execute the certain block of code repeatedly until test condition is false. Loops are used in performing repetitive task in programming.

There are 3 types of loops in C programming:

1. for loop
2. while loop
3. do...while loop

For Loop Syntax

for (initialization statement; test expression; update statement)

```
{  
    code/s to be executed;  
}
```

How for loop works in C programming?

The initialization statement is executed only once at the beginning of the for loop. Then the test expression is checked by the program. If the test expression is false, for loop is terminated. But if test expression is true then the code/s inside body of for loop is executed and then update expression is updated. This process repeats until test expression is false.

This flowchart describes the working of for loop in C programming.

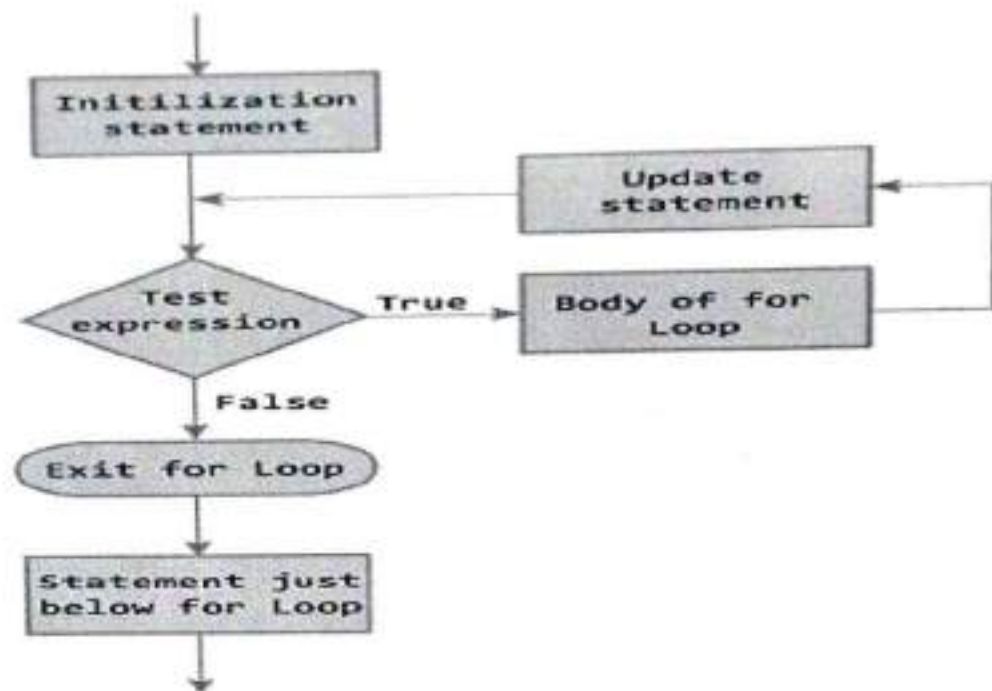


Figure: Flowchart of for Loop

11 (ii) Write a C-program to compute the sum of the series

/ C program to find sum of series $1+x+x^2+x^3+\dots+x^n$ */*

```

#include<stdio.h>
#include<conio.h>
#include<math.h>
void main()
{
  int i,n;
  float x,sum=0;
  clrscr(); //to clear the screen
  printf("1 + x + x^2 + x^3 + ..... + x^n ");
  printf("\n Enter the value of x and n:");
  scanf("%f %d",&x,&n);
  for(i=1;i<=n;++i)
  sum+=pow(x,i);
  sum++;
  printf("\nSum=%f",sum);
  getch(); //to stop the screen
}
  
```

VI Semester BSc 5 (CBCS) Degree examination,
 September/October - 2023
 Chemistry Paper - I (Regular)

I. 1) Limitations of CFT.

(i) The reason for considering all types of ligands as point charge is not satisfactorily explained by CFT.

(ii) As delocalisation of ligand electrons and orbital become more, there is increase in covalent bond which CFT do not take into account.

(iii) The CFT fails to explain the relative positions of ligands in the spectrochemical series like SCN^- is weaker ligand compared to NCS^- in series.

(iv) The theory does not explain as when there is formation of tetrahedral complex and square planar complex.

(v) The positions and intensities of spectral bands calculated on the basis of CFT are not always coincide with those determined experimentally.

(2) Small or trace quantities of elements or biomolecules required for living organisms are called as Trace elements.

Ex: Mo, Mn, Fe, Co, Cu, Zn etc

(3) Biological role of Sodium (Na)

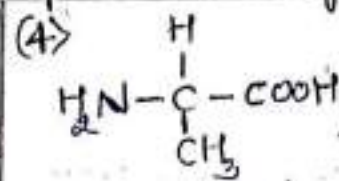
(i) Major cation in extracellular fluid

(ii) Osmotic balance

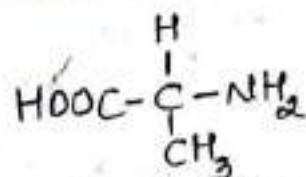
(iii) Charge carrier and electrolytic balance

(iv) Required in process of nerve impulse creation and its transmission.

(v) Involved in cotransport of sugars & Amino acids into the cells.



L-Alanine



D-Alanine

(5) Spontaneous change in specific rotation of an optical isomer until a constant value is reached is called as mutarotation.

(6) Epimers are diastereomers which differ in the configuration of only one carbon atom or asymmetric carbon atom.

Epimerisation - process of conversion of one epimer to another epimer.

(7) Induced dipole moment is the dipole moment which has occurred due to induction of charge by the action of the other molecule

or
 Induced dipole moment is a dipole moment created in a nonpolar molecule or neutral atom by coming into close contact with a polar molecule or charge particle such as an ion.

(8) The transition between orbitals having the same symmetry is said to be a symmetry forbidden transition
 Ex: molecules with a centre of symmetry shows symmetry forbidden transition.

(9) De-Broglie's hypothesis says that matter consists of both the particle nature as well as wave nature. De-Broglie wavelength is given as-

$$\lambda = \frac{h}{p} = \frac{h}{mv}$$

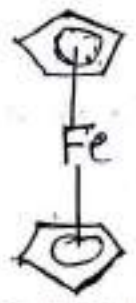
where p - particle momentum with mass m and velocity v
 h - Planck's constant

From above relation it can be said that the wavelength of matter is inversely proportional to the magnitude of particle linear momentum and is applicable to both microscopic and macroscopic particles.

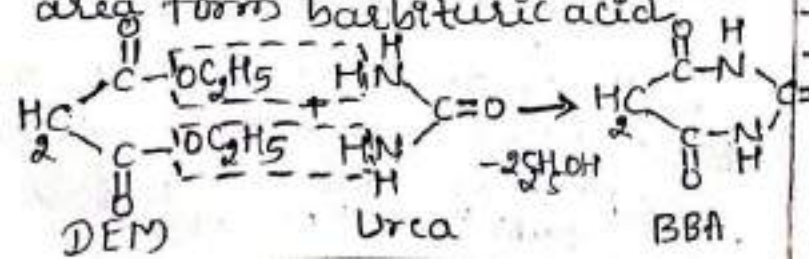
(10) Terpenes are the hydrocarbons of plant origin of the general formula $(C_5H_8)_n$ as well as their oxygenated, hydrogenated and dehydrogenated derivatives.

Ex: Citral, Myrcene, Zingiberene etc.

(11) Ferrocene $[Fe(C_5H_5)_2]$
 η^6 -bis(cyclopentadienyl) Iron (II)

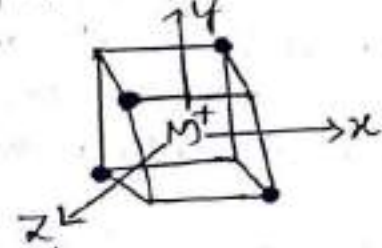


(12) Synthesis of Barbituric acid
 Malonic ester on condensation with urea forms barbituric acid



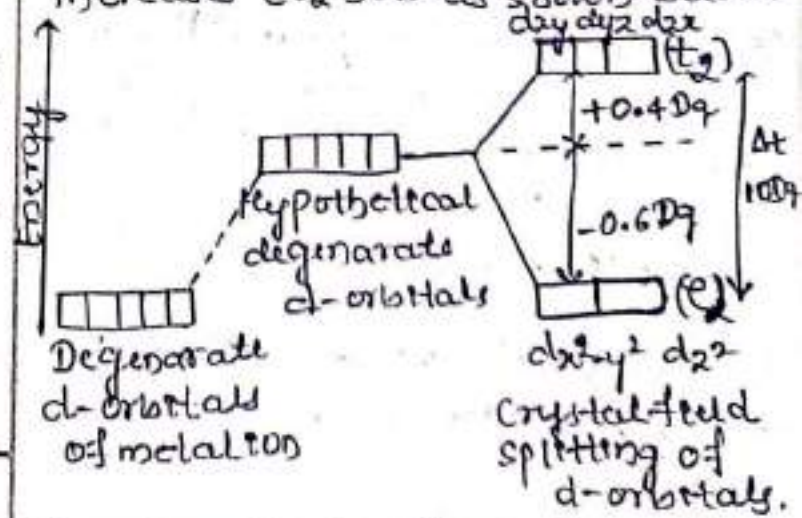
II. (1) Crystal field splitting of d-orbitals in tetrahedral complexes.

In tetrahedral complex, metal ion is at the centre of cube and four ligands occupy the four alternate corners of cube which constitutes a tetrahedron. The three axes are x, y & z as shown in figure



Ligands approach in between the axes towards the metal ion to form coordinate complex. The dx^2-y^2 and dz^2 orbitals lie along the axes while dxy, dyz & dxz orbitals lie in between the axes.

dx^2-y^2 & dz^2 orbitals experience less repulsion from ligand field and their energy decreases (e set). The energy of dxy, dyz & dxz orbitals increases (t_2 set) as shown below.



The magnitude of energy difference between the two sets of orbitals (e & t_2) is called as the crystal field splitting energy (CFSE)

i.e. $\Delta_t = 10Dq$
 -1 for tetrahedral

The energy of t_2 orbital is $+0.4 \Delta t$ higher than that of hypothetical degenerate d-orbitals and the energy of e orbital is $-0.6 \Delta t$ lower than that of hypothetical degenerate d-orbitals.

Each electron entering the t_2 orbital destabilizes the complex by $0.4 \Delta t$ and each electron entering the e orbital stabilizes by an amount of $0.6 \Delta t$.

The Ca^{+2} , Mn^{+2} and Zn^{+2} metal ions have d^0 , d^5 and d^{10} configurations in these cases d-orbitals are empty, half filled and completely filled respectively, these have zero CFSE because of spherical field around the nucleus.

In case of $\text{Sc}^{+2} (t_{2g}^1 e_g^0)$, $\text{Ti}^{+2} (t_{2g}^2 e_g^0)$ and $\text{V}^{+2} (t_{2g}^3 e_g^0)$, the electrons in d-orbitals occupy t_{2g} orbital of lower energy that is no electron present in e_g orbital and the electron from t_{2g} orbital which are oriented in between the orbital plane. Thus there is no shielding of nuclear charge and ligands are drawn closer to the nucleus. As a result its ionic radii is less than expected when there was no crystal field effect.

In case $\text{Cr}^{+2} (t_{2g}^3 e_g^1)$, the e_g orbitals has 1 electron, this orbital oriented in the direction of ligand hence it provides good shielding to the ligands and ligands are no longer approach so closely and ionic radii increases.

M^{+2} (high spin) $(t_{2g}^3 e_g^2)$ results in the spherical distribution of charge around the nucleus and the ionic radii is same as calculated.

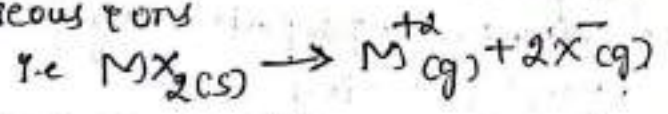
The $\text{Fe}^{+2} (t_{2g}^4 e_g^2)$, $\text{Co}^{+2} (t_{2g}^5 e_g^2)$ and $\text{Ni}^{+2} (t_{2g}^6 e_g^2)$ metal ionic radii decreases because the electron enters into e_g orbital.

$\text{Cu}^{+2} (t_{2g}^6 e_g^3)$ ionic radii is smaller than $\text{Zn}^{+2} (t_{2g}^6 e_g^4)$ where electrons are spherically distributed.

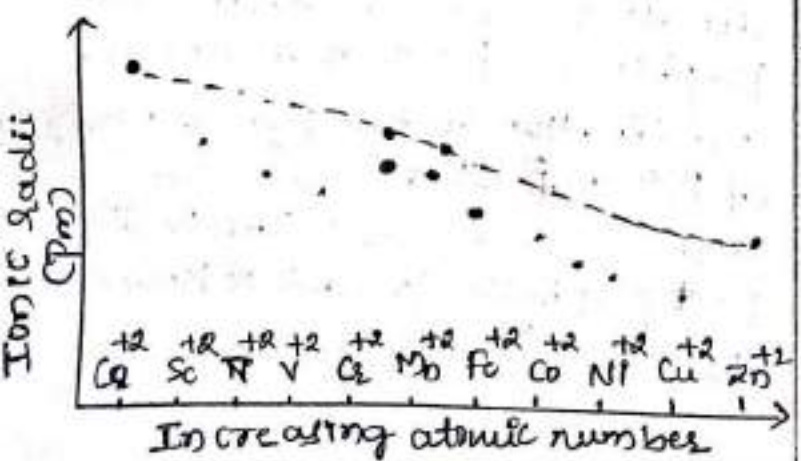
(2) Consequences of Crystal field splitting on ionic radii of M^{+2} ions.

The variation of ionic radii of bivalent cations of the metals of 1st transition series in their crystalline compounds in which each M^{+2} ion is octahedrally co-ordinated to six chlorides.

The lattice energy is the enthalpy change required to convert one mole of solid into its constituent gaseous ions



In the absence of CFSE (due to d-orbital splitting) the ionic radii for M^{+2} ions expected to decrease from d^1 to d^{10} .



(3) Colours of the complexes.

CFT can be used to explain the colours of transition metal complexes. A complex with 0 or 10 d electrons i.e. d^0 & d^{10} is colourless.

A complex must have partially filled d subshell on metal to exhibit colour. When the complex shows the colour it means it absorbs specific wavelengths of visible light having 400 - 700 nm. The colour observed is complementary colour of colour absorbed. Electrons absorb energy (wavelength) that can overcome Δ_0 value.

If the Δ_0 is large then more energy required and electron absorbs in short wavelength region and vice-versa. We know that Δ_0 value depends on metal & ligand i.e. when electron gets excited from lower d-orbital to higher d-orbital energy level on absorbing energy is known as d-d transition.

Magnetic properties of the complex

Based on the nature of electron either paired (Diamagnetic) or unpaired (Paramagnetic) the magnetic property arises.

The electron configuration of the transition metals changes when in a compound due to repulsive force between electrons in the ligands.

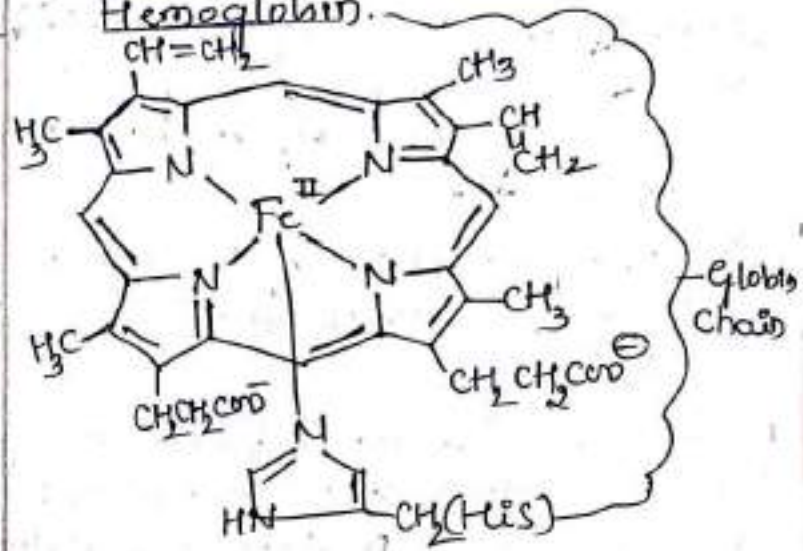
Based on nature of ligand i.e. strong field ligand decreases paramagnetic nature and increases number of unpaired electrons where as weak field ligand increases

Paramagnetic nature due to the unpairing of electrons

$$\mu = \sqrt{n(n+2)} \text{ BM.}$$

μ - magnetic moment
 n - no. of unpaired electrons.

(4) Structure and Function of Hemoglobin.



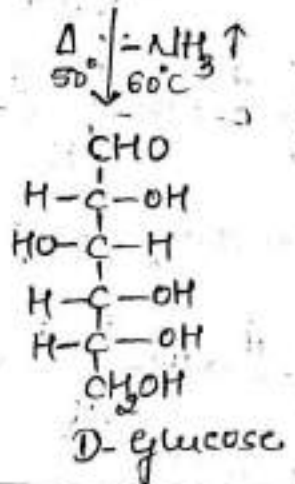
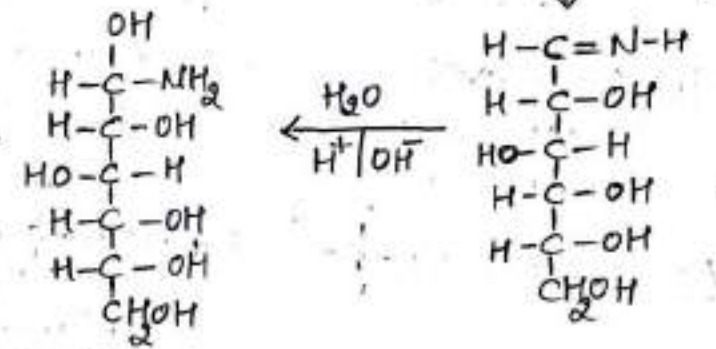
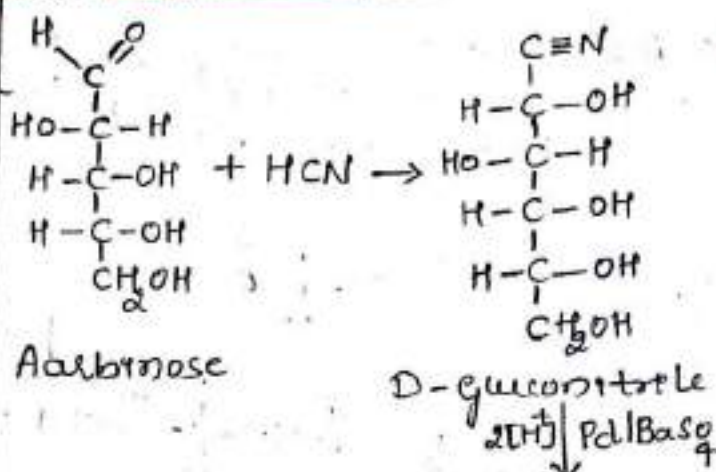
Fe-porphyrin referred to as heme is the prosthetic group of Hb & Mb. Heme derived from porphyrin. An unsubstituted tetrapyrrole connected at the α carbon by methylidene linkage to form porphyrin. Substitution on 8 pyrrole positions on a porphyrin ring produces a porphyrin ring. Different substitution on porphyrin ring at the eight pyrrole positions produce protoporphyrin IX, which is used in heme. The porphyrins can donate two protons from each porphyrin ring to form α -anions. Coordination by four pyrrole nitrogens of P IX to Fe^{II} produces an anionic heme unit. The Fe^{II} protoporphyrin IX unit is known as Heme b.

The 5th position is coordinated by the imidazole nitrogen of proximal histidine of globin chain.

Heme unit cannot itself carry O₂ but when it is folded with globin protein it can perform the task. In planar porphyrin ring due to presence of conjugated double bonds π & π* orbitals are available these allow the charge transfer electron transitions to give red colour of blood.

Functions:

- Transport of oxygen.
- electron transport.



III
(1) Killiani's - Fischer Synthesis.

→ It is a method for synthesizing monosaccharides. It involves the conversion of aldopentose to aldohexose. This synthesis is used for the extending or lengthening the monosaccharide chain by single carbon. There is a formation of stereogenic center there by epimers are formed.

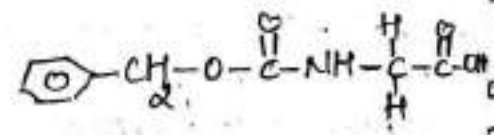
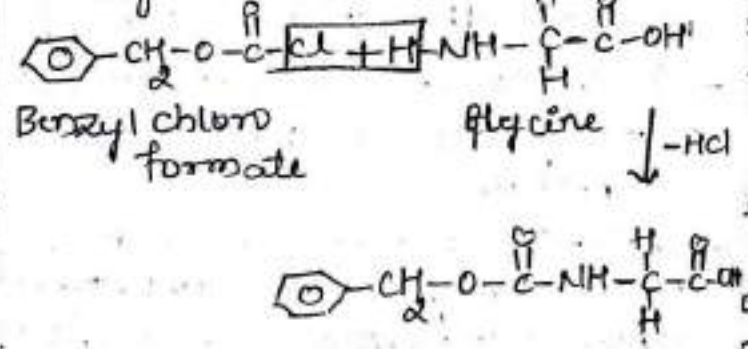
Aldopentose	→	Aldohexose
5 C compound		6 C compound
Arabinose		glucose

Reaction of arabinose with hydrogen cyanide to form cyanohydrin, in the reaction new stereogenic or asymmetric carbon is created. Hence a mixture of isomeric cyanohydrins is formed. Glucosamine is undergo reduction in the presence Pd/Baso it converts to amine followed by hydrolysis and heating loss of ammonia to form D-glucose.

(2) Bergmann synthesis of a dipeptide

This synthesis follows different steps as below.

(i) protection of the amino group of one molecule of α amino acid using benzyl chloro formate.



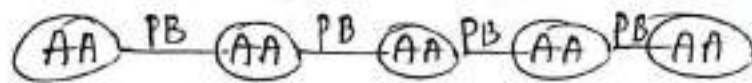
(A) Structure of Proteins.

The study of structure of complex proteins is undertaken through the following stages.

(i) Primary structure

This is concerned with the formation of chain of α -amino acids through the formation of peptide bonds.

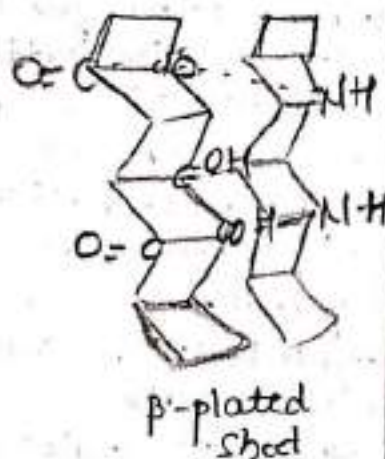
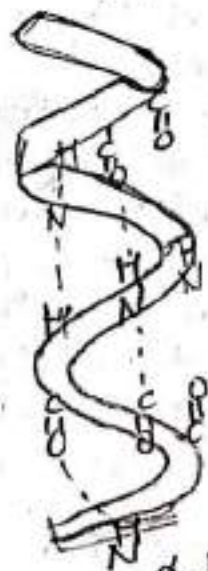
Proteins may have one or more polypeptide chain. Each polypeptide in a protein has amino acids linked with each other in a specific sequence is known as primary structure. It is immature protein because they where not provide any other function



(ii) Secondary structure

→ It refers to the shape in which along polypeptide chain can exist. They are found to exist in 2 different types α -helix and β -pleated sheet structure.

These structures arise due to the regular folding of the backbone of polypeptide chain due to hydrogen bonding between $-C=O$ & $-NH$ group of the peptide bond.

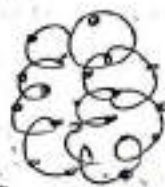


(iii) Tertiary structure

It represents overall folding of the polypeptide chain. i.e. further folding of 2° structure. This refers to a three dimensional structure of proteins. This is as a result of twisting, bending & folding of helices. It may be globular fibrous.

(iv) Quaternary structure

This proteins are composed from two or more polypeptide chains referred to as subunits. The spatial arrangement of these subunits with respect to each other is known as quaternary structure.



3° structure



4° structure

IV.

(1) Electronic transitions.

An electronic transition takes place when an electron moves from higher occupied MO (HOMO) to the lowest unoccupied MO (LUMO) by the absorption of suitable energy in the UV region.

Mainly there are four electronic transitions are there namely-

(i) $\sigma \rightarrow \sigma^*$ transition.

It is a high energy process since σ bonds are in general very strong. For saturated hydrocarbons like methane propane etc absorption occurs near 150nm (high energy). Consider $\sigma \rightarrow \sigma^*$ transition in a saturated hydrocarbon.

(ii) $n \rightarrow \sigma^*$ transition.

This type of transition takes place in saturated compounds containing one heteroatom with unshared pair of electrons (n electrons). Such transition requires less energy than that of

$\sigma \rightarrow \sigma^*$ transition.

Ex: Halides, ethers, alcohols, aldehydes, ketones, amines
water absorbs at 167 nm
Methyl alcohol absorbs at 174 nm
Methyl chloride at 169 nm.

In saturated alkyl halides, the energy required for such a transition decreases with increase in the size of the halogen atom or decrease in the electronegativity of atom
 $\text{CHI}_3 - 258 \text{ nm}$.

(iii) $\pi \rightarrow \pi^*$ transition.

This type of transition occurs in the unsaturated centres of molecule. i.e. in compounds containing double or triple bond & also in aromatics. The excitation of π electrons requires smaller energy & hence transition of this type occurs at longer wave length.

Ex: alkenes, alkynes, carbonyl compounds, cyanides etc.

(iv) $n \rightarrow \pi^*$ transition.

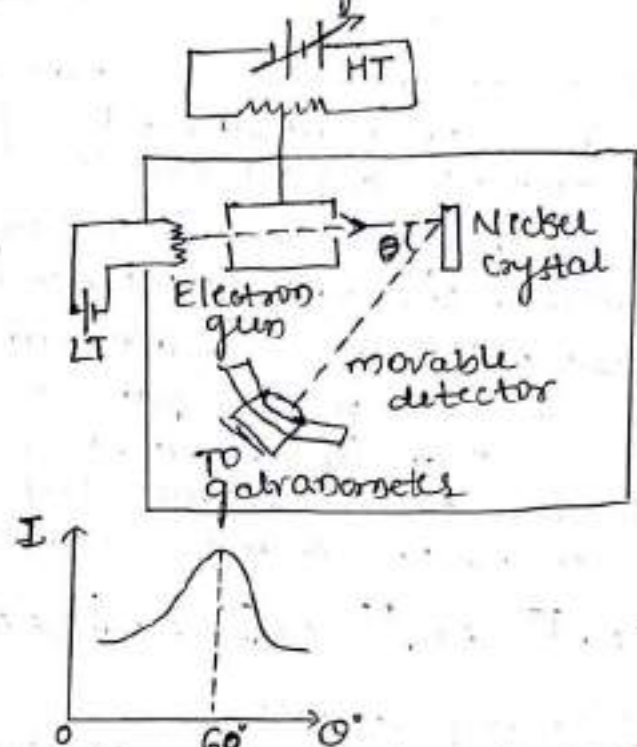
An electron of unshared electron pair on the heteroatom gets excited to π^* antibonding orbital. This type of transition requires least amount of energy out of all the transitions.

Ex: Saturated aldehydes
 $n \rightarrow \pi^*$ at 290 nm.

2) Experimental verification of DAVISSON-GERMER experiment.

Waves undergo properties like interference and diffraction, in order to prove electrons as waves, electron must also undergoes interference and diffraction.

In this experiment we will study about electron undergo diffraction when they are incident on a nickel crystal. In order to occur diffraction size of obstacle must be comparable with the wavelength of the wave.



From the de-Broglie hypothesis, it was found that wavelength of electron was 0.1 nm which is equal to distance between two nickel atoms therefore nickel crystal is used in this experiment. The experimental set up consists of -

(i) Electron gun which consists of tungsten filament coated by BeO as work function of barium oxide is very low, thus very little amount of heat is required to eject an electron when electron gun is heated by low tension (6V) battery, electrons are emitted, these emitted electrons are accelerated with the help of high tension battery, after that these electrons are passed through cylinder with a small hole & made to fall on nickel target, then electrons are scattered at different angles by the atoms of crystal. Movable detector connected with galvanometer will measure intensity of scattered electron.

This experiment was performed by varying the accelerating voltage from 44 V to 68 V. A strong peak was observed in the intensity of scattered electrons for $V = 54$ V at scattering angle $\theta = 60^\circ$.

This peak was observed due to constructive interference of scattered electrons from different layers of atoms of Nickel crystal.

From electron diffraction measurements wavelength of matter wave was found to be $\lambda = 0.165$ nm.

From theoretical de-Broglie's eqⁿ

$$\lambda = \frac{1.227}{\sqrt{V}} \text{ nm}$$

$$= \frac{1.227}{\sqrt{54}} \text{ nm}$$

$$\lambda = 0.167 \text{ nm}$$

Comparing the experimental and theoretical values they are in good agreement thus Davisson-Germer experiment confirms the wave nature of electrons.

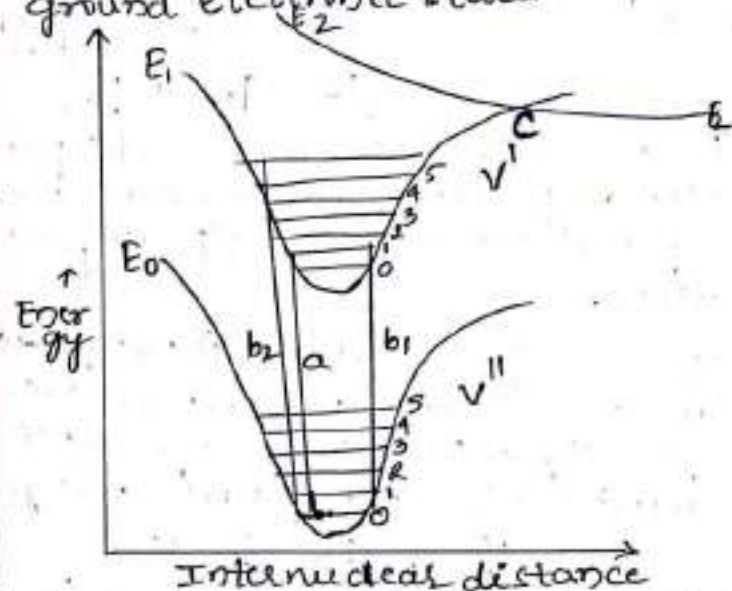
(3) Franck Condon Principle

It states that "an electronic transition takes place so rapidly that a vibrating molecule does not change its internuclear distance appreciably."

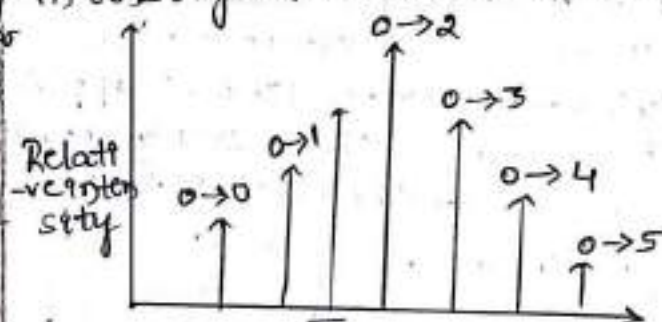
During electronic transition the nuclei do not change their position hence an electronic transition may be represented by a vertical line on a plot of potential energy versus the internuclear distance.

Consider a diatomic molecule, two potential energy curves for the molecule in the ground electronic state (E_0) & in the first electronic state (E_1). Since bonding in the excited state is weaker than in the ground state, the minimum in the potential energy curve for the excited

state occurs at a slightly greater internuclear distance than the corresponding minimum in the ground electronic state. Quantum mechanically it is known that the molecule is in the centre of the ground vibrational level of the ground electronic state.



When an energy falls on the molecule the most probable electronic transition (according to Franck Condon principle) takes place from $v''=0$ to $v'=2$ (i.e. $0 \rightarrow 2$) other transitions will take place with lower probability so that their relative intensities are smaller than the intensity of the $0 \rightarrow 2$ transition.



In molecules like H_2, N_2 etc (homonuclear) the highest occupied molecular orbital (HOMO) from ground state is a bonding molecular orbital and lowest unoccupied molecular orbital (LUMO) from excited ground state is an antibonding molecular orbital. Hence electron undergoes transition from (HOMO) to (LUMO).

(4) Determination of dipole moment by temperature variation method

→ This is based on Debye equation
Debye equation is -

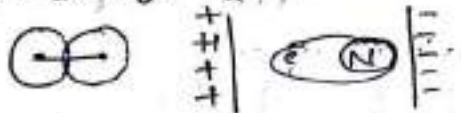
Molecular Polarisation (P_m) = Induced Polarisation + Orientation Polarisation

$$P_m = P_{in} + P_o$$

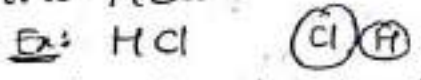
Polarisation: The process of separation of positive and negative charge by a distance d on applying electric field.

Induced polarisation: A non polar molecule is one in which centre of gravity of positive charge coincide with centre of negative charge

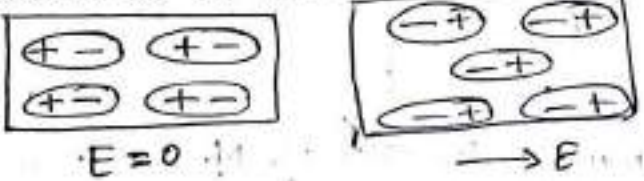
Ex: H₂, O₂, N₂, Cl₂ etc



Orientation polarisation: If molecules are permanent dipole they are oriented towards applied electric field



Polar molecules randomly oriented in absence of electric field. But when electric field is applied they oriented themselves in opposite direction of the electric field



$$P_m = P_{in} + P_o$$

$$P_m = \frac{4}{3} \pi N d + \frac{4}{3} \pi N \frac{\mu^2}{3kT}$$

$$P_m = a + \frac{b}{T} \quad \text{--- (1)}$$

a & b are constants

$$a = \frac{4}{3} \pi N d, \quad b = \frac{4 \pi N \mu^2}{9 \cdot 3k}$$

$$\mu = \sqrt{\frac{9k}{4 \pi N}} \cdot \sqrt{b}$$

where T - temperature

N - Avogadro Number

k - Boltzmann constant
 $1.380649 \times 10^{-23} \text{ J/K}$

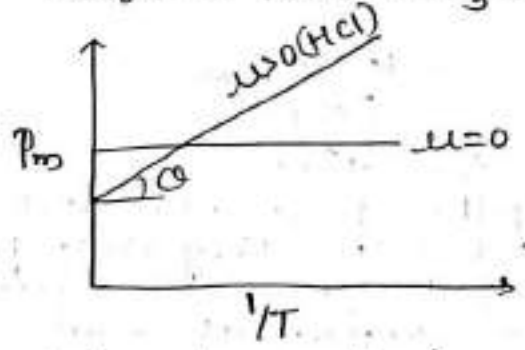
$$\mu = 0.0128 \times 10^{18} \sqrt{b} \quad \text{--- (2)}$$

According to eqⁿ (1) if a graph is plotted between P_m vs 1/T we get a straight line with slope equal to b. If straight line makes angle α with 1/T then α = b.

Substituting b value in eqⁿ (2) μ can be calculated

μ = 0 → CCl₄, CH₄, H₂, O₂ etc

μ ≠ 0 → HCl, CHCl₃, HF



Alternate method

$$P_{m1} = a + \frac{b}{T_1}$$

$$P_{m2} = a + \frac{b}{T_2}$$

$$P_{m1} - P_{m2} = a + \frac{b}{T_1} - a - \frac{b}{T_2}$$

$$P_{m1} - P_{m2} = b \left[\frac{1}{T_1} - \frac{1}{T_2} \right]$$

$$P_{m1} - P_{m2} = b \left[\frac{T_2 - T_1}{T_1 T_2} \right]$$

$$b = (P_{m1} - P_{m2}) \left(\frac{T_1 T_2}{T_2 - T_1} \right)$$

where P_{m1} - molar polarisation of molecule at temperature T₁

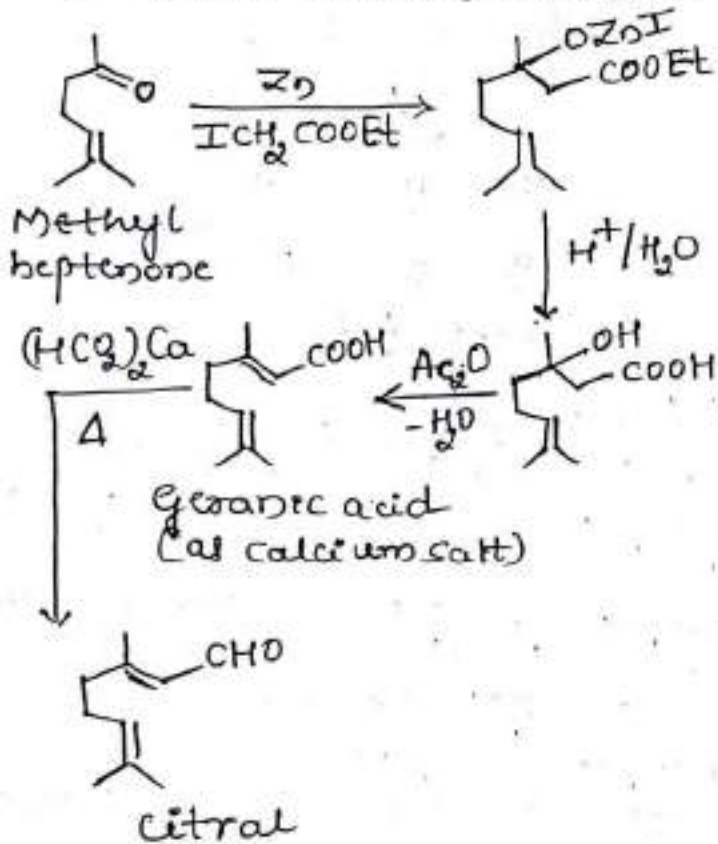
P_{m2} - molar polarisation of molecule at temperature T₂

V

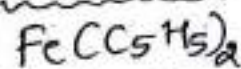
(1) Synthesis of Citral

The synthesis of citral from methyl heptenone involves various steps and it occurs through the following steps as shown below

- (i) Reformatsky reaction
- (ii) Hydrolysis with dilute acid
- (iii) Dehydration with acetic anhydride.
- (iv) distilling calcium geranate with calcium formate.

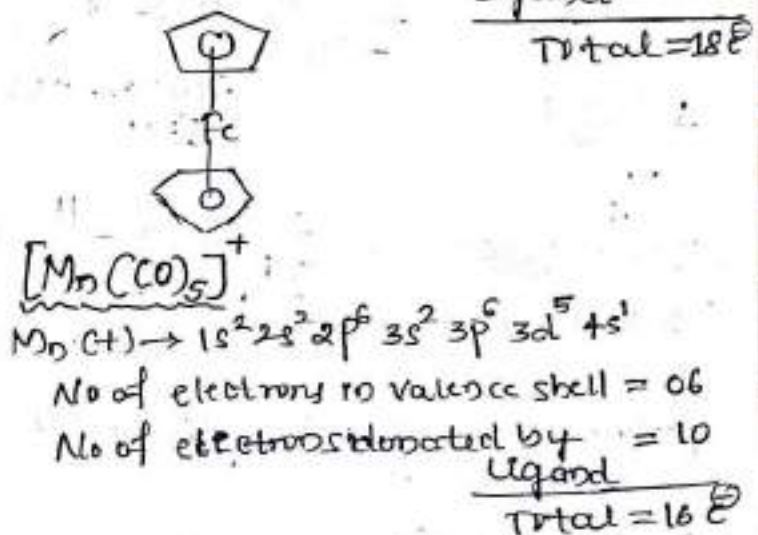
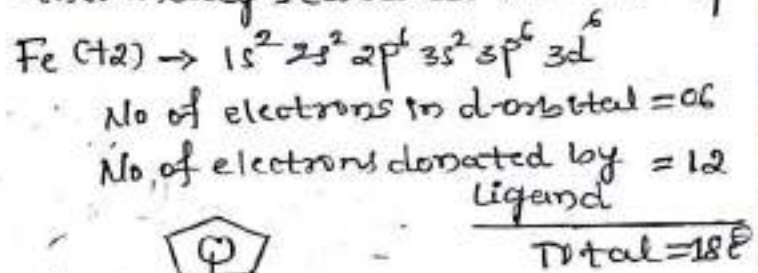


2) 18-electron rule with respect to Ferrocene.



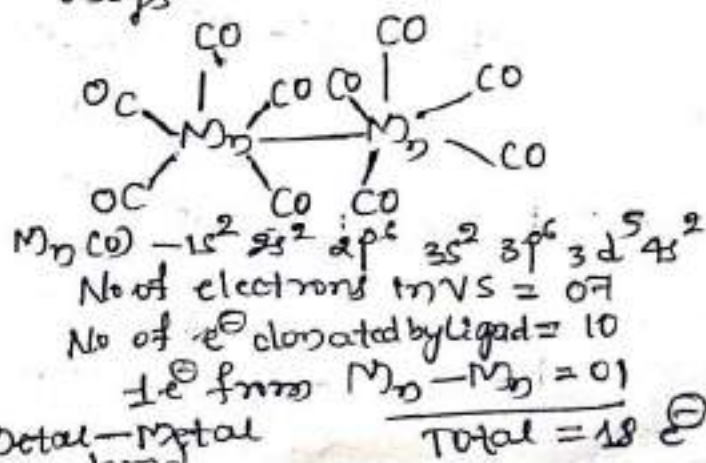
It is the prototypical metallocene a type of organometallic compound consisting of two cyclopentadienyl ring bound on opposite sides of a central metal atom, such organometallic compounds are also known as sandwich compounds.

In terms of bonding, the iron centre in ferrocene is usually assigned to the +2 oxidation state. Each cyclopentadienyl ring (Cp) is then allocated a single negative charge, bringing the net of π -e⁻ on each ring to six & thus making them aromatic. These 12 e⁻ (6 from each ring) are then shared with a metal via covalent bonding, when combined with the six d e⁻ on Fe⁺² the complex attains an 18-e⁻ configuration and becomes thermally stable at room temp.



Does not obey 18-e⁻ rule but $[\text{Mn}(\text{CO})_5]_2$ i.e. $[\text{Mn}_2(\text{CO})_{10}]$ dimanganese decacarbonyl $(\text{CO})_5\text{Mn}-\text{Mn}(\text{CO})_5$

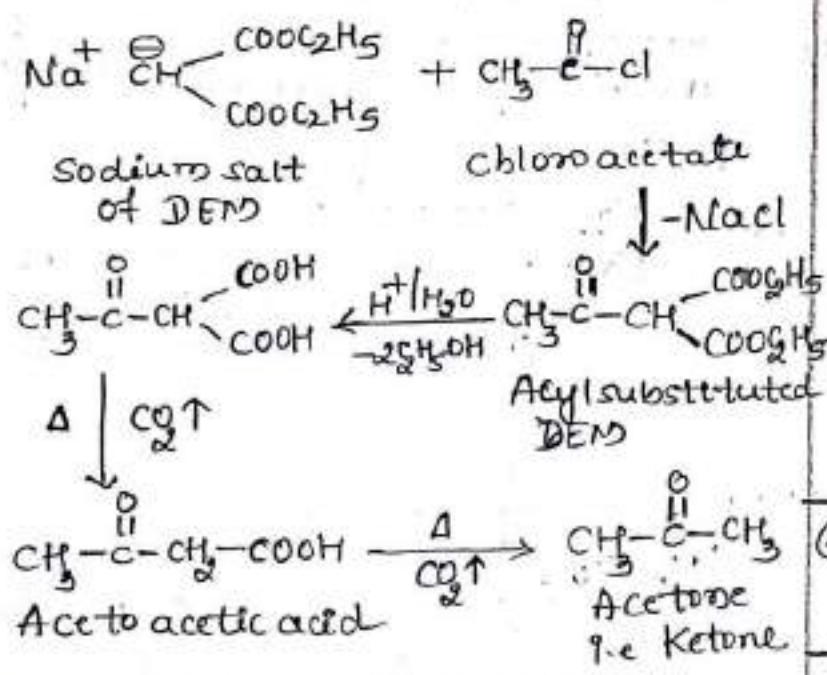
obeys



Metal-Metal

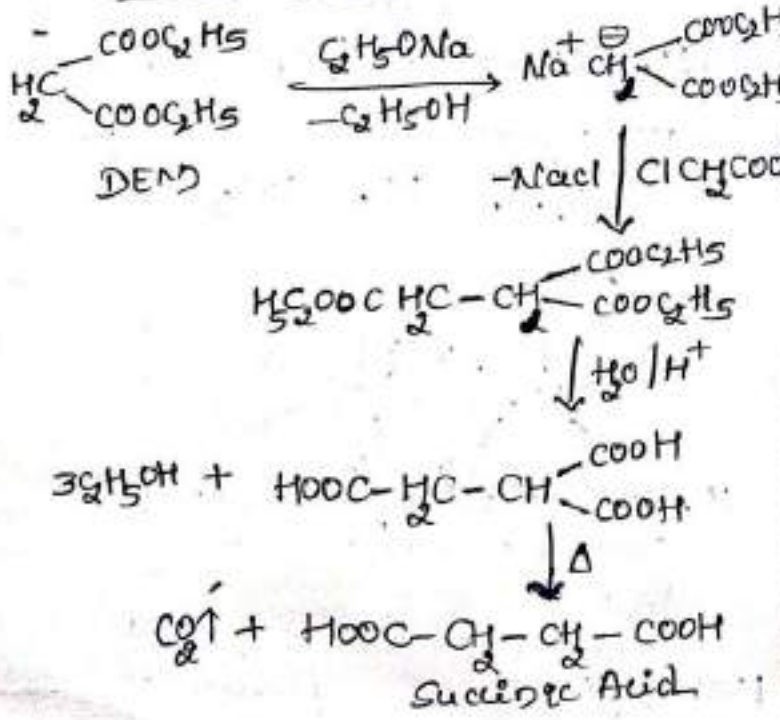
(3) Synthesis of Ketones from Diethylmalonate.

When sodium salt of diethyl malonate is treated with chloroacetate, acyl substituted diethyl malonate is formed followed by hydrolysis and double decarboxylation, simplest ketone i.e. acetone is formed.

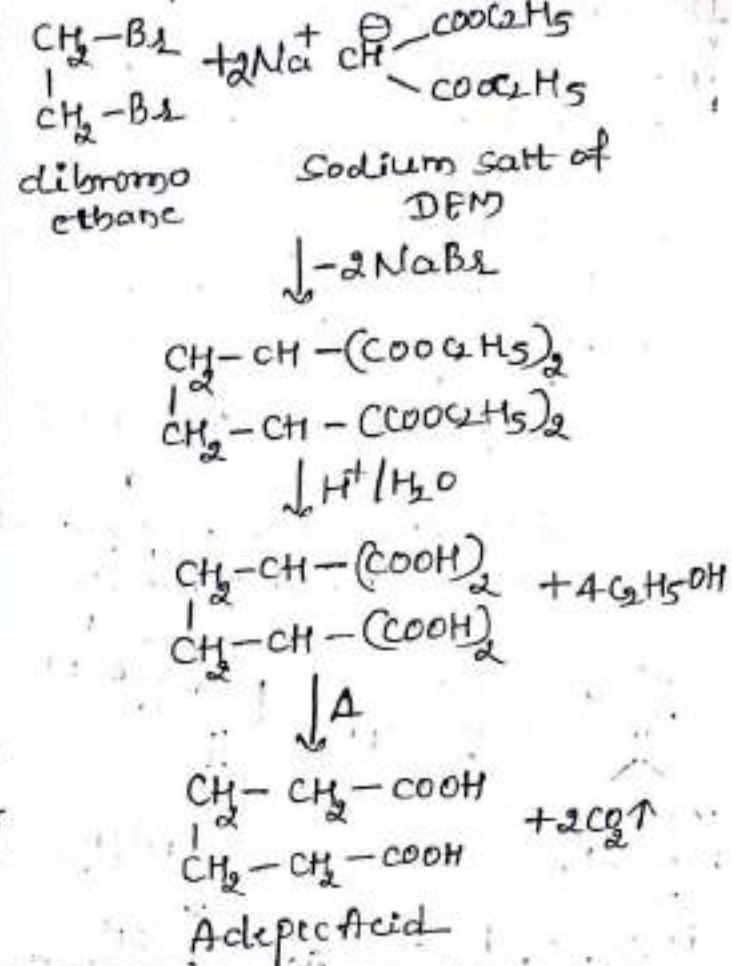


Synthesis of dicarboxylic acids from DEM

(1) Succinic Acid



(2) Adelic Acid



(4) Advantages of organic reagents over inorganic reagents.

- Organic reagents are more specific and selective than inorganic reagents.
- Organic precipitation is non ionic in nature and therefore it reduces co-precipitation in most cases.
- Organic reagents are very sensitive even a small quantity of substance can be detected.
- The complexes formed by organic reagents with metal ions are insoluble in water hence they can be easily separated, dried & weighed.
- Lesser amount of organic reagent is generally required for precipitation in gravimetric analysis hence it is more economical.
- ORs form coloured complexes with metal ions & intensity of colour changes with concⁿ of the metal ion & hence they can be estimated by colorimetric method.

QP CODE: 44961/F0210

VI Sem B.Com. 3(CBCS) Degree Exam,
Sept/Oct-2023

**PRINCIPLES AND PRACTICE OF AUDITING
SCHEME OF EVALUTION**

Dr. Vivekanand A. Mane (Mob: 8431209855)

Section-A

I. Answer any Ten of the following (10 X 2 =20)

- "Auditing is concerned with the verification of accounting data determining the accuracy and reliability of accounting statements and reports." - R.K. Mautz
- Audit note book is a diary or register maintained by audit staff to note errors, doubtful quarries and difficulties. It is used for recording important points to be included in the auditor's report.
- An audit programme is a detailed plan of the auditing work to be performed, specifying the procedure to be followed in verification of each item in the financial statements and giving the estimated time required.
- It is an arrangement of the duties of members of staff in such a manner that the work performed one person is automatically and independently checked by the other and the possibility of error or fraud is reduced to the minimum.
- It implies the audit of accounts by the staff of the business. Internal audit is an appraisal activity within an organization for the review of the accounting, financial and other operations as basis for protective and constructive service to the management.
- A written document that supports the entries made in the record books that helps in verifying the accuracy of the transactions.
- The report in which auditors express a qualified view of financial statements is a qualified audit report. It means that the company's financial records are not maintained in accordance with GAAP but no misinterpretations are involved.
- A contingent liability is a liability or a potential loss that may occur in the future depending on the outcome of a specific event. Potential lawsuits, product warranties, and pending investigation are some examples of contingent liability.
- Verification means the procedures normally carried out at the year end, to confirm the ownership, valuation and existence of items at the balance sheet date. In simple words verification means, 'proving the truth or conformation.
- (i) Institute of Chartered Accountants of India
(ii) Standards Auditing Practices
- Title of Report, Addressee details, Opening Paragraph, scope Paragraph, Opinion Paragraph, Signature, Place of Signature....
- Routine checking is a form of audit examination in which certain books and records that are common to all types of businesses are checked.

Section-B

II. Answer any Three of the following (3 X 5 =15)

- An audit that is performed after the financial period has ended and the accounts are prepared is known as a **final audit**. It continues in a session until all of the audit work has been completed. It is sometimes referred to as an annual or recurring audit.
(2 marks)

Advantages:

- It is not very expensive and more suitable for small concerns.
- It involves less time
- Less scope for alteration of figures.
- Convenient for Management.
- Thorough Checking.
(3 Marks)

3. Duties of an auditor:

- Provide an Audit Report
- Make Proper Enquiry
- Assist in Branch Audit
- Compliance with Auditing Standards
- Reporting of Frauds
- Provide Assistance in Investigation
- Adhere Principles of Auditing
- Provide Negative Opinion

4. Vouching of Credit Sales

- The sales register should be examined with copies of sales invoices. The sale of capital items should not be recorded in the sales book, otherwise the profits will be inflated.
- Test check should be applied on the calculations made in sale invoices.
- The totalling and the castings of sales book should be verified.
- Sales Tax, duties collected through sales invoices must be recorded under separate accounts.
- It should be verified that all sales invoices are prepared on the basis of challans and then sales invoices are entered in sales book and from there, posted to their respected accounts.
- Sales made in the current year must be recorded under that year and shall not be treated as sales of subsequent year.
- All cancelled sales invoices must be kept together for verification by auditor, who should see that cancelled invoices are properly treated in the books.
- The statement of accounts should be verified by getting confirmations from the customers

5. Auditing V/s Investigation

- An Audit is carried out to ensure that the balance Sheet and the profit and Loss A/C show a true and fair picture. But, on the other hand, an investigation is carried out on for some predefined purpose e.g. to know the financial position of the company or its earning capacity
- An Audit is limited only for an examination of the accounts of the concern but the

investigation covers not only examination of accounts, it involves probing deep into the matter and looking for required information far behind the books whenever necessary.

3. The Investigation is not legally compulsory but audit is statutorily compulsory in case of joint stock companies,
4. Auditing can only be conducted by a chartered accountant but it is not necessary that an investigator must be a Chartered accountant.
5. An audit is always carried out on behalf of the owner of the business, but the investigation may be conducted on behalf of the proprietor of the business, in case he suspects any fraud, or on behalf of the outside parties.
6. An audit always relates to a period of 1 year or 6 months but the investigation may cover several years.

6. Characteristics of audit report:

1. It is a medium through which auditor express his opinion on financial statements and condition of a business
2. Statement of facts collected by auditor.
3. Serves as medium to convey auditor's opinion.
4. Final or end product of audit.
5. Based on facts and information relating to company.
6. Essential for shareholders.
7. Report may be short or long.
8. Duly signed by auditor and attached to balance sheet.

Section-C

III. Answer any Two of the following (2X 15 =30)

7. **Auditor:** Auditors are qualified professionals who analyse firms' financial records to detect misrepresentations or discrepancies (2 marks)

Qualities of an auditor: (7 Marks)

1. The qualified chartered accountant can be appointed as auditor of a limited company.
2. The auditor must have thorough knowledge of principles and practice of all aspects of accountancy. He must be familiar with all systems of accountancy in use.
3. He should have adequate knowledge of financial management, industrial administration and business organization.
4. He must have thorough knowledge of audit case laws as per the various cases decide by the courts in and outside India.
5. He should be able to understand the technical details of business whose accounts he is going to audit.
6. An auditor must be honest i.e. He must certify that he does not believe to be true and he must take reasonable care and skill before he believes what he certifies is true.

7. He must act impartially and not influenced by others, directly or indirectly while discharging his duties.
8. He should be hard working, systematic and methodical.
9. He must have capacity to hear arguments of others,
10. He should have adequate skills and courage to write audit report correctly clearly and concisely.
11. He should not disclose the secrets of his client

Rights of an Auditor (6 Marks)

1. Right to Access books of accounts
2. Right to obtain the information and explanation
3. Right to receive notice
4. Right to sign audit report
5. Right to seek legal and technical advice
6. Right to remuneration
7. Right to be indemnified

8. Vouching: (3 x 5=15 Marks)

- a) Cash Sales
- b) Bills Payable
- c) Purchase of building

9. **Unqualified report:** When an auditor is satisfied with the fairness of the balance sheet and profit and loss account he will give a clean report. The auditor makes various statutory affirmations without reservations He is said to have the given an unqualified report on financial statements of the company. (3 marks)

Specimen: 12 marks

10. **Continuous audit** is an audit, which is conducted continuously throughout the year at regular or irregular intervals during the financial year. (3 Marks)

Periodical audit is an audit, which commences after closing of accounts and preparation of final accounts and is carried out continuously till the conclusion of audit.

Continuous audit: (6 Marks)

1. **Nature:** Book-keeping work and audit work is done simultaneously and after the transactions have been recorded in the books, these are checked together.
2. **Submission of audit report:** Audit report can be submitted quickly after the completion of the final account.
3. **Detection of errors and frauds:** Errors and frauds can be detected quickly and thus corrective action can be taken.
4. **Suitability:** It is suitable for big organization where transactions are too many.

Periodical audit: (6 Marks)

1. **Nature:** Audit works taken up only after all **book-keeping and accounting** work relating to the financial year have been completed.
2. **Submission of audit report:** Audit work takes a long time to be completed after the preparation of final accounts and so the submission of the audit reports also get delayed.
3. **Detection of errors and frauds:** As detailed examination of accounts is not possible in periodic audit, errors and frauds may remain undetected.
4. **Suitability:** It is suitable for small and medium sized organizations.

Section-C

IV. Case Study (Compulsory) (1X15=15)

11. Types of errors and Rectifications:

1. **Error of Commission:**
Entry is to be made as Credit Sales for Rs. 75,000/- Sales book
2. **Error of Principle**
Entry is to be taken in to respective assets account and balance sheet
3. **Error of Commission**
Entry is to be recorded in Genl. expenses/ Misc. expenses account
4. **Error of Commission**
Entry is to be recorded in Salary account
5. **Error of Omission**
Entry is to be recorded in Cash book

Marks Allocation:

Error: 2 marks each

Rectification: 1 marks each



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Theme
**'The Relevance of William Shakespeare
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This is to certify that *Prachi M. Jogale* of *B.K. College, Chikodi* has participated/ presented ~~his~~ her paper entitled *Shakespeare's Tragic Heroes* in A One-Day Seminar for Students on The Relevance of William Shakespeare in the 21st Century, held on 23rd April, 2024 at K.L.E. Society's Lingaraj College, Belagavi (Autonomous) and won.....place.


DR. SHASHIKANT KONNUR
HEAD, DEPARTMENT OF
ENGLISH


SMT VIBHA HEGDE
IQAC COORDINATOR


DR. H.S. MELINAMANI
PRINCIPAL

The Relevance of William Shakespeare in the 21st century

Register No :U15DM22A0065

Name :Prachi Jogale

Class : B.A 3 semester

College : KLE'S Basavaprabhu Kore Arts, Science and commerce College Chikodi.

E-mail : jogaleprachi2@gmail.com

Shakespeare's Tragic Heroes:A Comparative Analysis of characters such as Othello , King Lear , Romeo.

William Shakespeare

William Shakespeare was a renowned English poet, playwright, and actor born in 1564 in Stratford-upon-Avon. His birthday is most commonly celebrated on 23 April (see When was Shakespeare born), which is also believed to be the date he died in 1616.

Characteristics of a Tragic Hero

The tragic hero's tragic flaw, or hamartia, is a key element of the tragic hero's character that leads to their undoing. In Othello, it is his jealousy. In Hamlet, it is his inactivity. In Macbeth, it is his ambition. In King Lear, it is his pride. The stories that Shakespeare tells do not have happy endings and the central characters are all tragic heroes. They are all tragic heroes because they each possess a tragic flaw that leads to their eventual death.

01. Macbeth

02. Othello

03. Romeo

04. Hamlet

01.Macbeth - The Tragic Hero Consumed by Ambition

Macbeth is the titular character and tragic hero of William Shakespeare's Macbeth. His widespread appeal as a character is seen in the audience's ability to identify with the numerous human qualities and personal attributes that he possesses. Macbeth's fatal flaw is his ungovernable ambition and over-ambitious nature, which are a direct result of the witches' prophecies. These same prophecies are instrumental in Macbeth's decision to murder the king; a decision that will eventually lead to his tragic downfall.

02. Othello - The Tragic Hero Destroyed by Jealousy

In Shakespeare's "Othello," Othello is the eponymous tragic hero whose downfall is caused by his own jealousy. Othello is a flourishing man of high status and pride, and his fatal flaw is he is contrasted in each of the characters' minds and is what sparks his ambition and jealousy to grieve. This is best exemplified by his secretary, Iago, who pretends to be Othello's confidant in persuading him that his wife, Desdemona, is unfaithful – Othello allows Iago to convince him of this and it leads to his downfall.

03. Romeo - The Tragic Hero of Forbidden Love

Romeo is often considered a tragic hero. Why? He's a "young... and youth was often equated with inexperience and immaturity... He is a lover cursed by feelings too sudden and too strong to redirect when circumstances change. Alongside the Titans... a "pair of star-crossed lovers" who are doomed to death by their love and by familial hatred. Their deaths ultimately bring their respective families together, ending the violence and hatred.

04. Hamlet - The Tragic Hero Paralyzed by Indecision

Suggests that a tragic flaw should not make a character entirely good (as this clearly does). Nevertheless, in other respects, we adhere strictly to Aristotle's own principles. As a whole, Hamlet does not truly fit this traditional model. We see neither a lucky hero who is successful and happy until a small error in their judgement leads to downfall, nor do we see an essentially good man who makes a mistake through *eta*, or human error. However, perhaps Shakespeare did not intend that Hamlet should adhere to this model. On one reading, Hamlet's fate could be interpreted a sinister political commentary or as a comment on a society which enforced its monarchs to such a degree that any would have been bound to slip up eventually. Both these interpretations however seem less satisfactory than the tragic hero theory. The likelihood is that Hamlet (and indeed all Shakespeare's tragic heroes) were not meant to fulfil any theoretical model. Rather, they might have been intended to fulfil the expectations of the predominantly middle class early modern audience. This audience was not looking for a reflection of the traditional tragic hero, nor tailoring their taste to fit what they read in Aristotle's *Poetics*. Hamlet as a character is not supposed to be a specific model of a tragic hero but rather reflect a specific example of the flawed, middle class hero; one increasingly affluent and cogitative, with the potential for moral greatness yet ultimately flawed.

Conclusion

All Shakespearean tragic heroes have a remarkable capacity for suffering and enduring pain. Their experiences are commented on by certain Britons as being "life's bitterest sorrows". We consider the Bard to be the only person to have achieved the required mastery in crafting tragic hero characters. These characters are not simple to understand on account of their extremely complexity. There are several differences between these tragic heroes, and yet, they have enough commonalities that permits us to generalize their character traits.

Reference

Mateusz Brodowicz



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K.L.E. Society's Lingaraj College, Belagavi (Autonomous) and ~~was~~.....place.


DR. SHASHIKANT KONNUR
HEAD, DEPARTMENT OF
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SMT VIBHA HEGDE
IQAC COORDINATOR


DR. H.S. MELINAMANI
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RANI CHANNAMMA UNIVERSITY

Vidyasangama, Belagavi- 591156, Karnataka, India

Department of Post Graduate Studies and Research in Commerce
ONE DAY NATIONAL LEVEL CONFERENCE


ON


MSMEs AND ECONOMIC DEVELOPMENT – OPPORTUNITIES AND CHALLENGES

30th January, 2024

Certificate

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Conference on "MSMEs AND ECONOMIC DEVELOPMENT – OPPORTUNITIES AND
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Rani Channamma University, Belagavi, Karnataka on 30th January, 2024.


Dr. Sachindra G. R.
Organizing Secretary


Prof. S. B. Akash
Conference Chairman


Prof. H. Y. Kamble
Conference Director



RANI CHANNAMMA UNIVERSITY

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Department of Post Graduate Studies and Research in Commerce
ONE DAY NATIONAL LEVEL CONFERENCE

on

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30th January, 2024

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Organizing Secretary


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Conference Director



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President, RCUCMCTA

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Dr. Akash S.B.
Finance Officer
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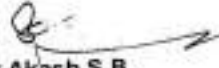
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
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
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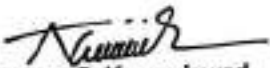
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
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SHRI SHIVAYOGI MURUGHENDRA SWAMIJI ARTS, SCIENCE
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This is to certify that ~~Mr./Ms~~ *Roopa. Desai*
OF *K.L.E's* *B.K* College *Chikodi*
has participated in State Level Quiz Competition on account of 'World Environment
Day' held on 05 June 2024 and Secured *I* Prize

[Signature]
COORDINATOR
SCIENCE ASSOCIATION

[Signature]
IQAC
COORDINATOR

[Signature]
CHIEF-GUEST

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PRINCIPAL



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This is to certify that **Mr./Ms** *Yallakka. Khot*
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[Signature]
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COORDINATOR

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CHIEF-GUEST

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K. L. E. Society's
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[Autonomous] Belagavi 590 001 (Estd.1935)

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State Level Poster Presentation Competition
On the eve of

WORLD ENVIRONMENT DAY

Organized by Department of Zoology In association with IQAC

CERTIFICATE OF APPRECIATION

This is to certify that Deepa B. Managanavi
of B. K. College, Chikodi has participated and presented in the
State Level Poster presentation competition organized by Department of Zoology on the occasion of
World Environment Day held on 5th June 2023.

Shri. Shidagoud S. Patil
Chairman, Zoological Association

Miss R. S. Maben
Head, Department of Zoology

Dr. M. M. Tamboli
IQAC Coordinator

Dr. Smt. J. S. Kawalekar
Principal



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State Level Poster Presentation Competition

On the eve of

WORLD ENVIRONMENT DAY

Organized by Department of Zoology In association with IQAC

CERTIFICATE OF APPRECIATION

This is to certify that Arpita . R. Naik
of B. K. College, Chikodi has participated and presented in the
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Shri. Shidagoud S. Patil
Chairman, Zoological Association

Miss R. S. Maben
Head, Department of Zoology

Dr. M. M. Tamboli
IQAC Coordinator

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Principal

K. L. F. Society's
Basavaprabhu Kore Arts, Science and Commerce College Chikodi
Department of Botany
B. Sc. I Semester Results JULY 2024 Exam Month-JAN-FEB 2024



Paper: Microbial diversity and Technology

Sl. No.	Reg. No.	Name of the students	Th.	In.	Pr.	In.	Total	%	Class
1	U15DM23S0017	Radhika A. Chougala	29	28	21	25	103	68.67	I class
2	U15DM23S0028	Umratasak L. Durakhe	34	37	20	23	114	76.00	Dist
3	U15DM23S0034	Bharatesh L. Kamate	24	36	19	23	102	68.00	I class
4	U15DM23S0048	Pallavi S. Badode	35	35	24	25	119	79.33	Dist
5	U15DM23S0050	Anupama A. Patil	36	32	25	25	118	78.67	Dist
6	U15DM23S0053	Kamaronissa A. Sayyad	39	35	21	24	119	79.33	Dist
7	U15DM23S0056	Madhuchandrika R. Sadalage	45	35	22	24	126	84.00	Dist
8	U15DM23S0076	Sakshi S. Kalal	25	31	17	22	95	63.33	I class
9	U15DM23S0082	Kamble Vishwajeet Anil	0	22	16	19	57	38.00	Fail
10	U15DM23S0084	Spoorti G. Dargi	31	36	20	20	107	71.33	Dist
11	U15DM23S0085	Suprema S. Jogalekar	29	31	22	24	106	70.67	Dist
12	U15DM23S0089	Bhoomika S. Chavan	Left college						
13	U15DM23S0090	Kavita S. Bendigeri	27	35	21	22	105	70.00	Dist
14	U15DM23S0093	Shweta B. Hundekar	33	36	18	24	111	74.00	Dist
15	U15DM23S0095	Basagouda C. Koti	15	29	17	17	78	52.00	II class
16	U15DM23S0097	Gayatri M. Mariyai	36	35	20	20	111	74.00	Dist
17	U15DM23S0104	Kaveri L. Boraganvi	28	36	22	22	108	72.00	Dist
18	U15DM23S0108	Prajwal P. Gadage	Left college						
19	U15DM23S0109	Vijay A. Mujawar	2	21	12	19	54	36.00	Fail
20	U15DM23S0111	Sudharani K. Gudodagi	46	37	20	24	127	84.67	Dist

Summary

Total Candidates	20
Total Appeared	20
Total passed	16
Distinction	12
I class	3
II Class	1
Fail	2
Absent	NIL
Pcentage	88.88

Toppers

Rank	Reg. No.	Name	Th	In	Pr	In	Total	%	Class
1	U15DM23S0111	Sudharani K. Gudodagi	46	37	20	24	127	84.67	Dist
2	U15DM23S0056	Madhuchandrika R. Sadalage	45	35	22	24	126	84.00	Dist
3	U15DM23S0048	Pallavi S.Badode	35	35	24	25	119	79.33	Dist
3	U15DM23S0053	Kamarunissa A. Sayyad	39	35	21	24	119	79.33	Dist

We are happy to present the BSC I semester result of Jan-Feb 2024. The results are 88.88%. Out of 20 students 2 students left college during semester, 12 students passed in Distinction class and 2 Students failed due to their negligence.


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Department of Botany
B. Sc. III Semester Results JULY 2024 Exam Month-JAN-FEB 2024



Paper: Plant Anatomy and Developmental Biology

Sl. No.	Reg. No.	Name of the students	Th.	In.	Pr.	In.	Total	%	Class
1	U15DM22S0011	Srushti S. Kamble	56	38	24	24	142	94.67	Dist
2	U15DM22S0016	Anand S. Kondaguli	42	30	21	24	117	78.00	Dist
3	U15DM22S0022	Nandini Chavan	59	38	24	24	145	96.67	Dist
4	U15DM22S0029	Prashant P. Bani	45	30	18	24	117	78.00	Dist
5	U15DM22S0040	Supriya Bane	49	36	24	24	133	88.67	Dist
6	U15DM22S0050	Nikhil T. Devamane	52	32	20	24	128	85.33	Dist
7	U15DM22S0051	Priya D. Patil	53	38	25	24	140	93.33	Dist
8	U15DM22S0052	Santosh Waghmare	54	38	25	24	141	94.00	Dist
9	U15DM22S0077	Shweta T. Varute	58	39	24	24	145	96.67	Dist

Summary

Total Candidates	9
Total Appeared	9
Total passed	9
Distinction	9
Absent	0
Fail	0
Percentage	100.00

Toppers

Rank	Reg. No.	Name	Th	In	Pr	In	Total	%	Class
1	U15DM22S0022	Nandini Chavan	59	38	24	24	145	96.67	Dist
1	U15DM22S0077	Shweta T. Varute	58	39	24	24	145	96.67	Dist
2	U15DM22S0011	Srushti S. Kamble	56	38	24	24	142	94.67	Dist
3	U15DM22S0052	Santosh Waghmare	54	38	25	24	141	94.00	Dist

We are extremely happy to present the BSC III semester result of Jan-Feb 2024. The results are outstanding and appreciable. Results are 100% and all the students are passed in Distinction class.

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Department of Botany
B. Sc. V Semester Results JULY 2024 Exam month-JAN-FEB 2024

Paper I: Plant Morphology and Taxonomy

Paper II: Genetics and Plant Breeding

Sl. No.	Reg. No.	Name of the students	Paper I					Paper II					I & II	%	Class
			Th.	In.	Pr.	In.	Total	Th.	In.	Pr.	In.	Total			
1	U15DM21S0018	Akshata M. kore	10	26	16	17	69	43	24	16	24	107	176	58.67	Fail
2	U15DM21S0026	Arpita R. Naik	15	28	22	24	89	35	24	22	24	105	194	64.67	Fail
3	U15DM21S0036	Deepa B. Managanvi	46	39	25	24	134	54	39	25	25	143	277	92.33	Dist
4	U15DM21S0043	Kavita G. Jadhav	33	38	24	25	120	48	38	25	25	136	256	85.33	Dist
5	U15DM21S0054	Madiha M. Mulla	28	33	22	24	107	46	30	21	24	121	228	76.00	Dist
6	U15DM21S0058	Mahantesh G. Devanagol	21	32	22	24	99	24	32	22	24	102	201	67.00	I class
7	U15DM21S0059	Meenakshi G. Naik	43	40	24	25	132	46	40	25	25	136	268	89.33	Dist
8	U15DM21S0068	Muzafar A. Kalsigar	33	29	19	19	100	24	30	17	24	95	195	65.00	I class
9	U15DM21S0069	Nagesh A. Kagawade	49	40	24	25	138	56	40	24	25	145	283	94.33	Dist
10	U15DM21S0070	Navyasrushti N. Darnbal	34	30	17	20	101	25	31	16	24	96	197	65.67	I class
11	U15DM21S0071	Namrata S. Chimmat	25	27	17	20	89	27	28	16	23	94	183	61.00	I class
12	U15DM21S0077	Poonam B. Malage	41	37	23	25	126	48	36	24	24	132	258	86.00	Dist
13	U15DM21S0081	Pratik A. Magadum	25	26	18	23	92	24	28	19	23	94	186	62.00	I class
14	U15DM21S0088	Rajashree M. Tirodkar	46	40	18	20	124	46	37	21	24	128	252	84.00	Dist
15	U15DM21S0094	Roopa R. Desai	46	40	22	25	133	50	37	24	24	135	268	89.33	Dist
16	U15DM21S0103	Sangeeta K. Bambilwade	52	39	22	25	138	53	35	24	25	137	275	91.67	Dist
17	U15DM21S0119	Sourabh A. Rendale	31	31	19	24	105	48	31	17	22	118	223	74.33	Dist
18	U15DM21S0134	Uttam S. Varute	25	29	19	24	97	41	31	17	22	111	208	69.33	I class
19	U15DM21S0137	Veera L. Musaguppi	57	40	21	25	143	47	37	22	24	130	273	91.00	Dist
20	U15DM21S0141	Vishal S. Aihole	24	20	12	12	68	31	19	16	22	88	156	52.00	II class
21	U15DM21S0142	Yallaka N. Khot	41	40	22	25	128	48	38	25	24	135	263	87.67	Dist
22	U15DM21S0145	Laxmi N. Pavappagol	39	28	19	24	110	12	25	22	24	83	193	64.33	I class

Summary

Total Candidates	22
Total Appeared	22
Total passed	20
Distinction	12
I class	7
II Class	1
Fail	2
Absent	NIL
Percentage	90.90

Toppers

Sl. No.	Reg. No.	Name of the students	Paper I					Paper II					I & II	%	Class
			Th.	In.	Pr.	In.	Total	Th.	In.	Pr.	In.	Total			
1	U15DM21S0069	Nagesh A. Kagawade	49	40	24	25	138	56	40	24	25	145	283	94.33	Dist
2	U15DM21S0036	Deepa B. Managarvi	46	39	25	24	134	54	39	25	25	143	277	92.33	Dist
3	U15DM21S0103	Sangeeta K. Bambalwade	52	39	22	25	138	53	35	24	25	137	275	91.67	Dist

We are extremely happy to present the BSC V semester result of Jan-Feb 2024. The results are outstanding and appreciable. Out of 22 students 12 students got distinction marks, 7 students got I class, 1 students got II class and 2 students failed, Department hopes they will clear those papers in forthcoming exams.


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Department of Botany
B. Sc. II Semester Results FEBRUARY 2024 Exam Month-AUG-SEP 2023

Paper: Diversity of Non-Flowering Plants

Sl. No.	Reg. No.	Name of the students	Th.	In.	Pr.	In.	Total	%	Class
1	U15DM22S0002	Kiran G. Kamble	0	21	11	12	44	29.33	Fail
2	U15DM22S0011	Srushti S. Kamble	33	40	20	19	112	74.67	Dist
3	U15DM22S0016	Anand S. Kondaguli	25	31	17	15	88	58.67	II clas
4	U15DM22S0022	Nandini Chavan	54	40	24	24	142	94.67	Dist
5	U15DM22S0029	Prashant P. Bani	41	31	23	20	115	76.67	Dist
6	U15DM22S0040	Supriya Bane	42	38	17	15	112	74.67	Dist
7	U15DM22S0050	Nikhil T. Devamane	36	38	23	24	121	80.67	Dist
8	U15DM22S0051	Priya D. Patil	42	39	25	24	130	86.67	Dist
9	U15DM22S0052	Santosh Waghmare	45	40	24	24	133	88.67	Dist
10	U15DM22S0077	Shweta T. Varute	38	39	25	24	126	84.00	Dist

Summary

Total Candidates	10
Total Appeared	10
Total passed	9
Distinction	8
I class	NIL
II Class	1
Pass Class	NIL
Absent	NIL
Fail	1
Pcentage	90.00

Toppers

Rank	Reg. No.	Name	Th	In	Pr	In	Total	%	Class
1	U15DM22S0022	Nandini Chavan	54	40	24	24	142	94.67	Dist
2	U15DM22S0052	Santosh Waghmare	45	40	24	24	133	88.67	Dist
3	U15DM22S0051	Priya D. Patil	42	39	25	24	130	86.67	Dist

We are extremely happy to present the BSC II sem result of FEBRUARY 2024. The results are outstanding and appreciable. Out of 10 students 08 students got distinction marks, one student Passed in II class and one student failed in the exam because he was an average student and also neglected the instructions given by the teachers .


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B. Sc. IV Semester Results FEBRUARY 2024 Exam month-AUG-SEP 2023

Paper: Ecology and Conservation Biology

Sl. No.	Reg. No.	Name of the students	Th.	In.	Pr.	In.	Total	%	Class
1	U15DM21S0018	Akshata M. kore	Result withheld						
2	U15DM21S0026	Arpita R. Naik	25	31	23	22	101	67.3	I Class
3	U15DM21S0036	Deepa B. Managanvi	43	40	24	25	132	88.0	Dist
4	U15DM21S0043	Kavita G. Jadhav	43	39	25	24	131	87.3	Dist
5	U15DM21S0054	Madiha M. Mulla	39	35	23	24	121	80.7	Dist
6	U15DM21S0058	Mahantesh G. Devanagol	35	34	23	25	117	78.0	Dist
7	U15DM21S0059	Meenakshi G. Naik	51	40	20	24	135	90.0	Dist
8	U15DM21S0068	Muzafar A. Kalaigar	30	19	20	24	93	62.0	I Class
9	U15DM21S0069	Nagesh A. Kagawade	52	40	24	24	140	93.3	Dist
10	U15DM21S0070	Navyasrushti N. Dambal	37	35	22	23	117	78.0	Dist
11	U15DM21S0071	Namrata S. Chimmat	25	27	21	23	96	64.0	I Class
12	U15DM21S0077	Poonam B. Malage	44	39	25	25	133	88.7	Dist
13	U15DM21S0081	Pratik A. Magadum	25	19	20	20	84	56.0	II Class
14	U15DM21S0088	Rajashree M. Tirodkar	46	40	24	25	135	90.0	Dist
15	U15DM21S0094	Roopa R. Desai	56	40	23	24	143	95.3	Dist
16	U15DM21S0103	Sangeeta K. Bambalwade	50	40	25	24	139	92.7	Dist
17	U15DM21S0119	Sourabh A. Rendale	45	24	21	18	108	72.0	Dist
18	U15DM21S0134	Uttam S. Varate	31	23	20	16	90	60.0	I Class
19	U15DM21S0137	Veena L. Musaguppi	52	40	25	22	139	92.7	Dist
20	U15DM21S0141	Vishal S. Aihole	25	18	20	16	79	52.7	II Class
21	U15DM21S0142	Yallaka N. Khot	48	40	25	25	138	92.0	Dist
22	U15DM21S0145	Laxmi N. Payappagol	33	34	22	20	109	72.7	Dist

Summary

Total Candidates	22
Total Appeared	22
Total passed	21
Distinction	12
I class	6
II Class	3
Result withheld	1
Absent	NIL
Ppercentage	100.00

Toppers

Toppers									
1	U15DM21S0094	Roopa R. Desai	56	40	23	24	143	95.3	Dist
2	U15DM21S0069	Nagesh A. Kagawade	52	40	24	24	140	93.3	Dist
3	U15DM21S0103	Sangeeta K. Bambalwade	50	40	25	24	139	92.7	Dist
3	U15DM21S0137	Veena L. Musaguppi	52	40	25	22	139	92.7	Dist
4	U15DM21S0142	Yallaka N. Khot	48	40	25	25	138	92.0	Dist
5	U15DM21S0059	Meenakshi G. Naik	51	40	20	24	135	90.0	Dist
5	U15DM21S0088	Rajashree M. Tirodkar	46	40	24	25	135	90.0	Dist

We are extremely happy to present the BSC III sem result of December 2023. The results are outstanding and appreciable. Out of 22 students 12 students got distinction marks, 6 students got I class, 3 students got II class and 1 student Result is pending department hopes they will score good marks and thus we declare B.Sc. III sem results are 100%.


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Department of Zoology
B.Sc IV Semester DSC Results 2022-23
Result Declared by RCU on 10.02.2024



S. No	Exam Number	Name of the candidate	Theory (100)	Practical (50)	Total (150)	%	Result
1	U15DM21S0009	Aditya A. Nilajyoti	70	46	116	77.33	Distinction
2	U15DM21S0024	Ankita G. Vaddar	60	38	98	65.33	First class
3	U15DM21S0026	Arpita R. Naik	40	40	80	53.33	Fail
4	U15DM21S0031	Bhimrav A. Sajane	87	50	137	91.33	Distinction
5	U15DM21S0036	Deepa B. Managanvi	94	49	143	95.33	Distinction
6	U15DM21S0044	Kavita L. Mecchannavar	71	39	110	73.33	Distinction
7	U15DM21S0054	Madiha M. Mulla	64	42	106	70.67	Distinction
8	U15DM21S0058	Mahantesh G. Devanagol	56	45	101	67.33	First class
9	U15DM21S0069	Nagesh A. Kagawade	88	49	137	91.33	Distinction
10	U15DM21S0070	Navyasrushti N. Dambal	69	35	104	69.33	First class
11	U15DM21S0071	Namrata S. Chimmatt	56	34	90	60.00	First class
12	U15DM21S0076	Pawar R. Ritesh	59	38	97	64.67	First class
13	U15DM21S0077	Poonam B. Malage	82	46	128	85.33	Distinction
14	U15DM21S0091	Rohan S. Kamble	39	32	71	47.33	Fail
15	U15DM21S0094	Roopa R. Desai	92	48	140	93.33	Distinction
16	U15DM21S0096	Rutuja R. Chonchannavar	82	45	127	84.67	Distinction
17	U15DM21S0114	Sneha H. Kambar	59	45	104	69.33	First class
18	U15DM21S0137	Veena L. Musaguppi	81	50	131	87.33	Distinction
19	U15DM21S0139	Vishal A. Vadavade	67	39	106	70.67	Distinction
20	U15DM21S0141	Vishal S. Aihole	29	31	60	40.00	Fail
21	U15DM21S0142	Yallakka N. Khot	84	49	133	88.67	Distinction
22	U15DM21S0145	Laxmi N. Payappagol	53	40	93	62.00	First class

Total Number of student Appeared	22
Total Number of student Passed	3
% of Passing	86.36

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M.COM TOPPERS FOR THE YEAR 2022-23

Sl No	Name of the Student	Marks Obtained	Percentage	Rank
1	Miss. Shubhangi Naik	1791	74.62	1 st
2	Miss. Soumya Kumbar	1747	72.79	2 nd
3	Mr. Abhishek Naik	1742	72.58	3 rd


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Post Graduate Department of Botany

Result analysis of II Semester M.Sc. Examination: October (2023-2024)

Total Appeared – 05

Total Pass – 04

Result Withheld – 01

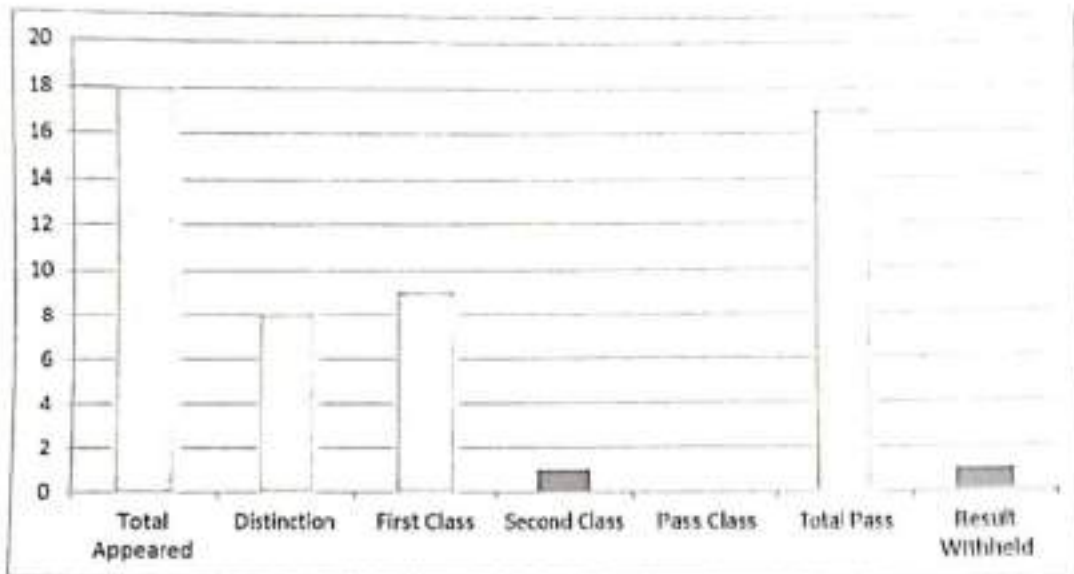
Percentage of passing –80%

Subject	Total Appeared	Distinction	First Class	Second Class	Pass Class	Total Pass	Result Withheld	% of Passing
Overall	05	02	02	-	-	04	01	80
2.1- Biochemistry and Biophysic	05	02	-	02	-	04	01	80
2.2- Developmental Biology	05	01	-	04	-	05	-	100
2.3- Genetics and plant Breeding	05	02	02	01	-	05	-	100
2.4- OEC (Economic Zoology)	05	02	02	01	-	05	-	100
Practical III- based on 2.1	05	03	02	-	-	05	-	100
Practical IV- Based on 2.2 and 2.3	05	03	02	-	-	05	-	100

Toppers:

Sl. No.	Name of the student	Marks obtained	SGPA
1	Vanita Balikayi	434	7.50
2	Prashant Uppar	419	7.17
3	Shweta Ashok Koli	413	6.42

Analysis:



- The results shows that, 02 students are represented with Distinction Grade (7.50 and 7.17) and 02 students represented with A Grade (6.42 and 6.33), four are passed and result withheld is 1.
- In this batch, **Miss. Vanita Balilayi** has scored 7.50 SGPA grade, secured first rank for the college. This is followed by **Mr. Prashant Uppar** with 7.17 SGPA grade and **Miss. Shweta Ashok Koli** with 6.42 SGPA grade.
- Total number of female 04 and male 01
- All students have studied well for the examination and secured good marks.

The Co-ordinator,
P.G. Department of Botany

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Post Graduate Department of Botany

Result analysis of IV Semester M.Sc. Examination

October 2023 (2023-2024)

Total Appeared – 18

Total Pass – 17

Result Withheld – 01

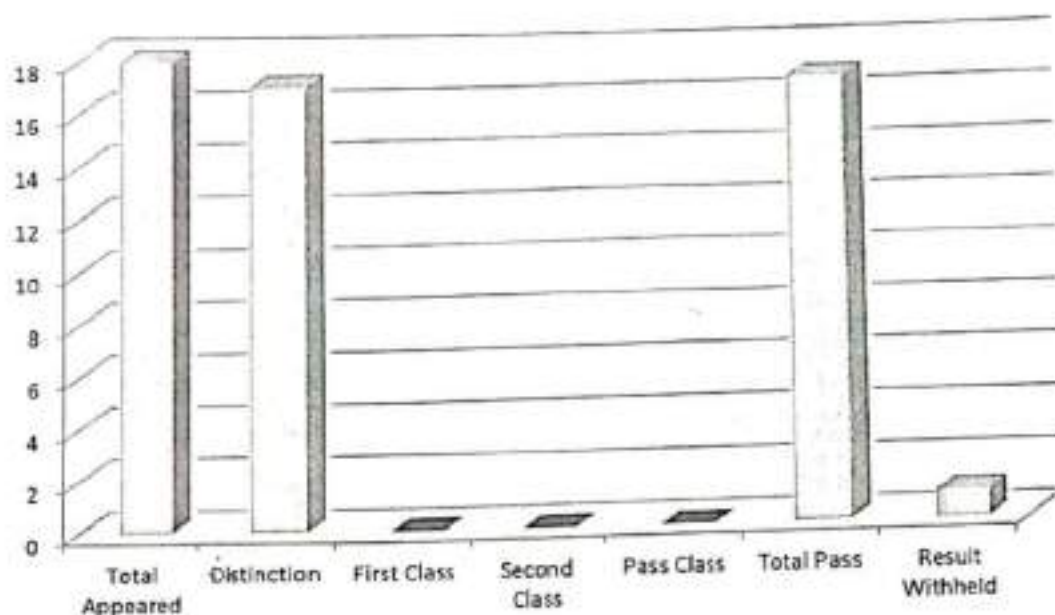
Percentage of passing – 94.44 %

Subject	Total Appeared	Distinction	First Class	Second Class	Pass Class	Total Pass	Result Withheld	% of Passing
Overall	18	15	2	-	-	-	1	94.44
4.1- Mycology and Plant Pathology	18	09	06	02	-	-	1	94.44
4.2- Ecology and Environmental Biology	18	16	1	1	-	-	-	100
4.3-Plant Biotechnology	18	14	2	1	-	-	1	94.44
4.4- Research methodologies and Techniques in Botany	18	13	5	-	-	-	-	100
Practical VII: Based on 4.1 and 4.2	18	18	-	-	-	-	-	100
Project	18	18	-	-	-	-	-	100

Toppers:


Sl. No.	Name of the student	Marks obtained	SGPA
1	Sudharani Sanadhi	475	8.33
2	Shweta Tirth	486	8.25
3	Akshata Kulkarni	482	8.08

Analysis:



- The result shows that, all the 15 students are represented with Distinction Grade (7.00 to 10.00), 02 students represented with A Grade (6.00 to 7.50), 17 student are passed and result withheld is 1.
- In this batch, Miss. Sudharani Sanadhi has scored 8.33 SGPA grade, secured first rank for the college. This is followed by Miss Shweta Tirth with 8.25 SGPA grade and Miss. Akshata Kulkarni with 8.08 SGPA grade.
- Total number of female 16 and male 02
- All students have studied well for the examination and secured good marks.


The Co-ordinator,
P.G. Department of Botany


Principal
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DEPARTMENT OF MATHEMATICS

Result Analysis B.Sc I- Sem (DSC) Jan/Feb 2024

Distinction	First Class	Second Class	Pass	Fail	Total Appeared	Total Pass	%
18	08	04	05	16	51	35	68.62

Top Scorer

1. Jyoti S. Thakkannavar 97/100
2. Priya L. Patil 91/100
3. Tejashwini A. Jodatti 89/100
4. Sangeeta B. Chougala 89/100
5. Aishwarya A. Torase 88/100
6. Jyoti K. Inchalakaranji 88/100
7. Arpita K. Holeppagol 87/100
8. Swati S. Hindalakar 86/100
9. Sudarshan S. Munnolli 85/100


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DEPARTMENT OF MATHEMATICS

Result Analysis B.Sc II- Sem (DSC) Aug/Sep 2023

Distinction	First Class	Second Class	Pass	Fail	Total Appeared	Total Pass	%
19	02	03	08	01	33	32	96.96

Top Scorer

1. Brahmi S. Patil 100/100
2. Shifa Mujawar 96/100
3. Srushti R. Savadi 93/100
4. Komal Gangale 90/100
5. Pradnya B. Ghali 88/100
6. Jyoti A. Champu 87/100
7. Amruta Chavan 85/100
8. Pallavi Patil 85/100
9. Bhagyashree Malage 85/100

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DEPARTMENT OF MATHEMATICS

Result Analysis B.Sc II- Sem (OEC) Aug/Sep 2023

Distinction	First Class	Second Class	Pass	Fail	Total Appeared	Total Pass	%
04	02	04	02	04	16	12	75

Top Scorer

1. Sneha S. Bidre 82/100
2. Megha V. Joshi 81/100
3. Priyanka A. Mali 80/100


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DEPARTMENT OF MATHEMATICS

Result Analysis B.Sc III- Sem (DSC) Jan/Feb 2024

Distinction	First Class	Second Class	Pass	Fail	Total Appeared	Total Pass	%
10	09	06	02	06	33	27	81.81

Top Scorer

1. Brahmi S. Patil 100/100
2. Amruta Chavan 87/100
3. Ranjita S. Koli 83/100
4. Pradnya B. Ghali 83/100
5. Vidya K. Kumbar 82/100
6. Jyoti A. Champu 80/100

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DEPARTMENT OF MATHEMATICS

Result Analysis B.Sc III- Sem (OEC) Jan/Feb 2024

Distinction	First Class	Second Class	Pass	Fail	Total Appeared	Total Pass	%
02	06	04	03	01	16	15	93.75

Top Scorer

1. Sneha S. Bidre

87/100

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DEPARTMENT OF MATHEMATICS

Result Analysis B.Sc IV- Sem Aug/Sep 2023

Distinction	First Class	Second Class	Pass	Fail	Total Appeared	Total Pass	%
22	10	07	08	05	52	47	90.38

Top Scorer

1. Daneshwari Kolalagi 97/100
2. Poornima Kagawade 97/100
3. Sandhya Sambaji 96/100
4. Shraddha Mutagi 96/100
5. Preeti Patil 95/100
6. Madhumati Meghannavar 94/100
7. Roopali Walake 93/100
8. Shravan Kattikar 90/100
9. Priya Jantennavar 88/100
10. Kumar Hikadi 86/100

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DEPARTMENT OF MATHEMATICS

Result Analysis B.Sc V- Sem Jan/Feb 2024

(Paper-I)

Distinction	First Class	Second Class	Pass	Fail	Total Appeared	Total Pass	%
19	10	11	05	04	49	45	91.84

Top Scorer

1. Daneshwari Kolalagi 100/100
2. Poornima Kagawade 100/100
3. Roopali Walake 95/100
4. Shraddha Mutagi 95/100
5. Niharika Banavanna 95/100
6. Madhumati Meghannavar 93/100
7. Shravan Kattikar 92/100
8. Ashwini Banage 85/100
9. Megha Havaladar 85/100

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DEPARTMENT OF MATHEMATICS

Result Analysis B.Sc V- Sem Jan/Feb 2024

(Paper-II)

Distinction	First Class	Second Class	Pass	Fail	Total Appeared	Total Pass	%
17	14	06	03	09	49	40	81.63

Top Scorer

1. Daneshwari Kolalagi 100/100
2. Shraddha Mutagi 98/100
3. Roopali Walake 95/100
4. Madhumati Meghannavar 89/100
5. Sandhya Sambhaji 87/100
6. Shravan Kattikar 85/100


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**Result Analysis B.Sc VI- Sem Sep/Oct 2023
(Paper-I)**

Distinction	First Class	Second Class	Pass	Fail	Total Appeared	Total Pass	%
44	25	15	21	02	107	105	98.13

Top Scorer

1. Nikita K. Danoli 100/100
2. Akshata A. Kamble 99/100
3. Shruti D. Korabu 98/100
4. Mahammadarif A. Dange 97/100
5. Keerti V. Kottalagi 96/100
6. Shreni R. Shiraganvve 95/100
7. Swapna A. Mangavate 95/100
8. Reshma R. Vernekar 91/100
9. Ashruta R. Ninganure 90/100
10. Nikhita S. Panadare 89/100

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DEPARTMENT OF MATHEMATICS

**Result Analysis B. Sc VI- Sem Sep/Oct 2023
(Paper-II)**

Distinction	First Class	Second Class	Pass	Fail	Total Appeared	Total Pass	%
59	26	15	06	01	107	106	99.06

Top Scorer

1. Nikita K. Danoli 100/100
2. Akshata A. Kamble 100/100
3. Mahammadarif A. Dange 100/100
4. Pragati M. Paranjape 100/100
5. Reshma R. Vernekar 100/100
6. Shruti D. Korabu 100/100
7. Kaveri C. Kamble 98/100
8. Bhagyashree G. Dooganavar 97/100
9. Swapna A. Mangavate 97/100
10. Mahananda S. Shipure 96/100
11. Shreni R. Shiraganvve 95/100

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DEPARTMENT OF MATHEMATICS

**Result Analysis B.Sc VI- Sem Sep/Oct 2023
(SEC)**

Distinction	First Class	Second Class	Pass	Fail	Total Appeared	Total Pass	%
20	01	--	01	--	22	22	100

Top Scorer

- | | |
|--------------------------|-------|
| 1. Nikita S. Panadare | 50/50 |
| 2. Pallavi K. Khot | 50/50 |
| 3. Pragati M. Paranjape | 50/50 |
| 4. Shruti D. Korabu | 50/50 |
| 5. Kaveri C. Kamble | 49/50 |
| 6. Ashruta R. Ninganure | 48/50 |
| 7. Priyanka B. Bamanale | 48/50 |
| 8. Suresh Chigari | 47/50 |
| 9. Nayana M. Drakshe | 46/50 |
| 10. Preeti T. Babhagouda | 46/50 |
| 11. Ramesh A. Patil | 46/50 |
| 12. Siddant K. Aski | 45/50 |

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DEPARTMENT OF MATHEMATICS

List of Centum Score 2023-24

Sl. No.	Name	Class	Paper
1	Miss. Nikita Danoli	B. Sc Sixth Sem	Paper I & II
2	Miss. Akshata A.Kamble	B. Sc Sixth Sem	Paper II
3	Mr. Mahammadarif A. Dange	B. Sc Sixth Sem	Paper II
4	Miss. Shruti D. Korbu	B. Sc Sixth Sem	Paper II
5	Miss. Pragati M. Paranjape	B. Sc Sixth Sem	Paper II
6	Miss. Reshma R. Vernekar	B. Sc Sixth Sem	Paper II
7	Miss. Daneshwari Kolalagi	B.Sc Fifth Sem	Paper I & II
8	Miss. Poornima N. Kagawade	B.Sc Fifth Sem	Paper I
9	Miss. Roopali Walake	B.Sc Fifth Sem	Paper II
10	Miss. Bramhi S. Patil	B.Sc Third	O D E


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Phone No.: 0831-2565237

ಫೋನ್ ನಂ. : ೦೮೩೧-೨೫೬೫೨೩೭

EXAMINATION SECTION**ಪರೀಕ್ಷಾಂಗ ವಿಭಾಗ**ವೆಬ್ ಸೈಟ್ : www.rcub.ac.inಈ-ಮೇಲ್ : registrareval@rcub.ac.in

ಕ್ರ. ಸಂ: ರಾಚವಿ/ಬೆಳಗಾವಿ/ಪ.ವಿ/2024-25/೨೨೨೯

ದಿನಾಂಕ: ೦೩-೦೯-೨೦೨೪

PROVISIONAL RANK LIST

The Provisional Rank List of B.Sc course for the Academic Year 2022-23 has been prepared. The Principals are requested to go through the Provisional rank holders list of September/October-2023 Examination. If any discrepancies are found you are hereby informed to bring it to the notice of The Registrar Evaluation on or before 09-09-2024.

Sl No	Reg. No	Name	Att.	Max. Marks	Sec. Marks	Percentage	Rank No	Caste	College Code	College Name
1	S2039638	KUMARESH NAGALINGAPPA KATARAKI	F/01	4900	4720	96.33	1	GM	6201	S V P S'S SRI. S B M ARTS, SCIENCE AND COMMERCE COLLEGE, BADAMI
2	S2029349	TAMLUL MADRASI	F/01	4900	4695	95.82	2	IIB	5206	ANJUMAN ARTS, SCIENCE AND COMMERCE COLLEGE, VIJAYAPURA
3	S2014451	MINAL BALESH BADIGER	F/01	4900	4650	94.90	3	IIA	4204	J E S SHRI. K A LOKAPUR A/S/C COLLEGE, ATHANI
4	S2016913	TEJASWINI NETAJI PATIL	F/01	4900	4647	94.84	4	IIIB	4222	DMSM'S B K COLLEGE OF ARTS, SCIENCE AND COMMERCE, BELAGAVI
5	S2040250	BHAGYASHREE DALAWAI	F/01	4900	4646	94.82	5	IIA	6204	BASAVESHWAR SCIENCE COLLEGE, BAGALKOT
6	S2018346	SHRUTI DAYANAND KORBU	F/01	4900	4632	94.53	6	IIIB	4251	B K COLLEGE OF ARTS, SCIENCE & COMMERCE, CHIKKODI
7	S2016270	SHAMBHAVI ANANT GHORPADE	F/01	4900	4614	94.16	7	IIIB	4218	SKES'S G S SCIENCE COLLEGE, BELAGAVI
8	S2040404	ROOPA NINGAPPA METI	F/01	4900	4592	93.71	8	IIA	6204	BASAVESHWAR SCIENCE COLLEGE, BAGALKOT
9	S2028785	SATISH NAGAPPAGOL	F/01	4900	4584	93.55	9	SC	5204	S B ARTS & KCP SCIENCE COLLEGE, VIJAYAPURA
10	S2044907	SHWETA SHREESHAIL THAKKANAVAR	F/01	4900	4584	93.55	9	IIIB	6260	TUNGAL SCHOOL OF BASIC & APPLIED SCIENCES, JAMKHANDI
11	S2044820	BHAGYA TIPPANNA TELI	F/01	4900	4565	93.16	10	IIA	6260	TUNGAL SCHOOL OF BASIC & APPLIED SCIENCES, JAMKHANDI

Sd/-

Registrar (Evaluation)

Rani Channamma University, Belagavi

Copy To,

1. Registrar, Rani Channamma University, Belagavi.
2. Finance Officer, Rani Channamma University, Belagavi.
3. PS Vice-Chancellor, Rani Channamma University, Belagavi.
4. The Principals of all Colleges affiliated to Rani Channamma University, Belagavi
5. Web Site Copy
6. Office Copy



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KARNATAKA EXAMINATIONS AUTHORITY



KARNATAKA STATE ELIGIBILITY TEST FOR ASSISTANT PROFESSOR

Certificate.No : 1064

(Valid in the State of KARNATAKA only)



K-SET APP.NO : KS1074979

K-SET Ref.No : 20231740134



Certified that Mr./Ms. PARVATI BEERAPPA IMMADI
Son / Daughter of BEERAPPA
and NIRAMAL
has qualified the K-SET for Eligibility for Assistant Professor held on **13th January 2024**
in the subject COMMERCE.

As per the information provided by the candidate, he / she had completed / appeared or was pursuing his / her Master's degree or equivalent examination in the related subject at the time of applying for K-SET.

The date of Eligibility for Assistant Professor is the date of declaration of K-SET result, i.e., **28th May 2024**, or the date of completion of Master's degree or equivalent examination with required percentage of marks within two years from the date of declaration of K-SET result, i.e., by **27th May 2026**, whichever is later.

The authenticity of this certificate and category in which the candidate had appeared/qualified should be verified from the Karnataka Examinations Authority (KEA) by the institution / appointing authority. The certificate can also be verified by scanning the QR Code printed on this certificate.

The validity of this certificate is forever.

Date of Issue : 13 JUL 2024


Prasanna H., IAS

Executive Director, KEA &
Member Secretary, K-SET




Prof. B.Thimme Gowda

Chairman, K-SET

NOTE:

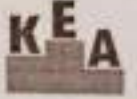
Karnataka Examinations Authority, has issued the certificate on the basis of information provided by the candidate in his/her Application Form. The appointing authority should verify the original records / certificates of the candidate while considering him/her for appointment, as Karnataka Examinations Authority, is not responsible for the same. The candidate must fulfill the minimum eligibility conditions for K-SET as laid down in the notification for the said K-SET.



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ಕರ್ನಾಟಕ ಪರೀಕ್ಷಾ ಪ್ರಾಧಿಕಾರ

Candidate Copy

Government of Karnataka
Karnataka Examinations Authority



K-SET 2023 DOCUMENTS VERIFICATION ACKNOWLEDGEMENT

SL.NO. : 6320

APP.NO. : KS1026044

CANDIDATE NAME: PRASHANT UPPAR

SUBJECT : LIFE SCIENCE(VII)



SL.NO.	Documnets List	Status
1	K-SET Admission Ticket.	Y
2	Marks Card of All Semesters.	N
3	Category/ Income Cerificate (For SC/ST in Form-D,Cat-1 in Form-E and 2A,2B,3A & 3B in Form-F Issued by the concerned Tahsildar of karnataka).	Y
4	Aadhar card Copy.	Y

I heareby declare that the information given by me is true to the best of my knowledge and I have received K-SET Cerificate from Karnataka Examinations Aurtority.

Date of Recivied :krei

Prashant Uppar

Signature of the Candidate

20/07/2024

[Signature]

Signature of the Verifier